

Chapter 2: Basic Chemistry

I. Elements and Atoms

- A. All matter is composed of a limited number of substances called **elements**.
- B. Each element is made up of building blocks called **atoms**.
- C. Atoms contain 3 major types of subatomic particles:
 - 1. Protons: **(p⁺)** charge
 - 2. Neutrons: **(n⁰)** charge
 - 3. Electrons: **(e⁻)** charge
- D. Particles with the same charge **repel** each other.
- E. Particles with unlike (**+/-**) charges **attract** each other.
- F. Elements are pure substances
 - 1. Elements cannot be broken down into simpler substances.
 - 2. List a few elements that are important to the human body: **oxygen, carbon, hydrogen, and nitrogen**

II. Chemical Bonds

- A. Ionic Bonds (See Figure 2.6)
 - 1. Ionic bonds are created when **electrons** are completely transferred from one atom to another. Ionic bonds form between ions.
 - a. Ions are **charged** particles.
 - i. Ions with a positive charge are **cations** (Example: **K⁺, H⁺**)
 - ii. Ions with a negative charge are **anions** (Example: **Cl⁻, Br⁻**)
 - 2. Example a compound formed by an ionic bond important in the human body: **H⁺, Na⁺**
- B. Covalent Bonds (See Figure 2.7)
 - 1. A bond in which two atoms **sharing** electrons
 - a. Nonpolar covalent bonds: atoms in this type of covalent bond share electrons
 - i. Example: **H₂, O₂**
 - b. Polar covalent bonds: atoms in this type of covalent bond **un-**equally share electrons. Such bonding forms a polar molecule because one end has a slight **negative** charge and the other end thus carries a slight **positive** charge
 - i. Example: **H₂O**
- C. Hydrogen bonds: a weak attraction between a slight positive charge on a **hydrogen** atom of one polar covalent bond and a weak negative charge on an **oxygen** or nitrogen atom of another polar covalent bond. (See Figure 2.9)
 - 1. Hydrogen bonding is very common between **water** molecules.

Chapter 2 Basic Chemistry Packet Answers

JG Myers



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