

## Chapter 20 Worksheets: Redox ANSWERS

1. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

- $2\text{Sr} + \text{O}_2 \longrightarrow 2\text{SrO}$   $\text{Sr}^0 \text{ to } \text{Sr}^{2+}$ ; oxidized/reducing agent  $\text{O}^0 \text{ to } \text{O}^{2-}$ ; reduced/ox. ag.
- $2\text{Li} + \text{S} \longrightarrow \text{Li}_2\text{S}$   $\text{Li}^0 \text{ to } \text{Li}^{+}$ ; oxidized/ox. ag.  $\text{S}^0 \text{ to } \text{S}^{2-}$ ; reduced/ox. ag.
- $2\text{Cu} + \text{Br}_2 \longrightarrow 2\text{CuBr}$   $\text{Cu}^0 \text{ to } \text{Cu}^{+}$ ; oxidized/ox. ag.  $\text{Br}^0 \text{ to } \text{Br}^{-}$ ; reduced/ox. ag.
- $3\text{Mg} + \text{N}_2 \longrightarrow \text{Mg}_3\text{N}_2$   $\text{Mg}^0 \text{ to } \text{Mg}^{2+}$ ; oxidized/ox. ag.  $\text{N}^0 \text{ to } \text{N}^{3-}$ ; reduced/ox. ag.
- $4\text{Fe} + 3\text{O}_2 \longrightarrow 2\text{Fe}_2\text{O}_3$   $\text{Fe}^0 \text{ to } \text{Fe}^{3+}$ ; oxidized/ox. ag.  $\text{O}^0 \text{ to } \text{O}^{2-}$ ; reduced/ox. ag.
- $\text{Cl}_2 + 2\text{NaBr} \longrightarrow 2\text{NaCl} + \text{Br}_2$   $\text{Cl}^0 \text{ to } \text{Cl}^{-}$ ; reduced/ox. ag.  $\text{Br}^{-} \text{ to } \text{Br}^0$ ; oxidized/ox. ag.
- $\text{Si} + 2\text{F}_2 \longrightarrow \text{SiF}_4$   $\text{Si}^0 \text{ to } \text{Si}^{4+}$ ; oxidized/ox. ag.  $\text{F}^0 \text{ to } \text{F}^{-}$ ; reduced/ox. ag.
- $\text{Mg} + 2\text{HCl} \longrightarrow \text{MgCl}_2 + \text{H}_2$   $\text{Mg}^0 \text{ to } \text{Mg}^{2+}$ ; oxidized/ox. ag.  $\text{H}^{+} \text{ to } \text{H}^0$ ; reduced/a.
- $2\text{Na} + 2\text{H}_2\text{O} \longrightarrow 2\text{NaOH} + \text{H}_2$   $\text{Na}^0 \text{ to } \text{Na}^{+}$ ; oxidized/a.  $\text{H}^{+} \text{ to } \text{H}^0$ ; reduced/a.

11. Give the oxidation number of each kind of atom or ion.

- |            |       |                    |                     |                     |            |             |
|------------|-------|--------------------|---------------------|---------------------|------------|-------------|
| a. sulfate | b. Sn | c. $\text{S}^{2-}$ | d. $\text{Fe}^{3+}$ | e. $\text{Sn}^{4+}$ | f. nitrate | g. ammonium |
| 2-         | 0     | 2-                 | 3+                  | 4+                  | 1-         | 1+          |

12. Calculate the oxidation number of chromium in each of the following.

- |                            |                                       |                    |             |               |
|----------------------------|---------------------------------------|--------------------|-------------|---------------|
| a. $\text{Cr}_2\text{O}_3$ | b. $\text{Na}_2\text{Cr}_2\text{O}_7$ | c. $\text{CrSO}_4$ | d. chromate | e. dichromate |
| 3+                         | 6+                                    | 2+                 | 7+          | 6+            |

13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations)

- $\text{C} + \text{H}_2\text{SO}_4 \longrightarrow \text{CO}_2 + \text{SO}_2 + \text{H}_2\text{O}$   $\text{C}^0 \text{ to } \text{C}^{4+}$ ; oxidized  $\text{S}^{6+} \text{ to } \text{S}^{4+}$ ; reduced
- $\text{HNO}_3 + \text{HI} \longrightarrow \text{NO} + \text{I}_2 + \text{H}_2\text{O}$   $\text{N}^{5+} \text{ to } \text{N}^{2+}$ ; reduced  $\text{I}^{-} \text{ to } \text{I}^0$ ; oxidized
- $\text{KMnO}_4 + \text{HCl} \longrightarrow \text{MnCl}_2 + \text{Cl}_2 + \text{H}_2\text{O} + \text{KCl}$   $\text{Mn}^{7+} \text{ to } \text{Mn}^{2+}$ ; reduced  $\text{Cl}^{-} \text{ to } \text{Cl}^0$ ; oxidized
- $\text{Sn} + \text{HNO}_3 \longrightarrow \text{SnO}_2 + \text{NO} + \text{H}_2\text{O}$   $\text{Sn}^0 \text{ to } \text{Sn}^{4+}$ ; oxidized  $\text{N}^{5+} \text{ to } \text{N}^{2+}$ ; red.

14. For each reaction in problem 13, identify the oxidizing agent and reducing agent.

- |                               |                          |
|-------------------------------|--------------------------|
| a. oxidizing agent: sulfur    | reducing agent: carbon   |
| b. oxidizing agent: nitrogen  | reducing agent: iodine   |
| c. oxidizing agent: manganate | reducing agent: chloride |
| d. oxidizing agent: nitrogen  | reducing agent: stannous |

# Chapter 20 Oxidation Reduction Reactions Worksheet

**Cracolice**

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### **Table of Contents Chapter 20 Oxidation Reduction Reactions Worksheet**

1. Understanding the eBook Chapter 20 Oxidation Reduction Reactions Worksheet
  - The Rise of Digital Reading Chapter 20 Oxidation Reduction Reactions Worksheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 20 Oxidation Reduction Reactions Worksheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 20 Oxidation Reduction Reactions Worksheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 20 Oxidation Reduction Reactions Worksheet
  - Personalized Recommendations
  - Chapter 20 Oxidation Reduction Reactions Worksheet User Reviews and Ratings
  - Chapter 20 Oxidation Reduction Reactions Worksheet and Bestseller Lists
5. Accessing Chapter 20 Oxidation Reduction Reactions Worksheet Free and Paid eBooks
  - Chapter 20 Oxidation Reduction Reactions Worksheet Public Domain eBooks
  - Chapter 20 Oxidation Reduction Reactions Worksheet eBook Subscription Services
  - Chapter 20 Oxidation Reduction Reactions Worksheet Budget-Friendly Options

6. Navigating Chapter 20 Oxidation Reduction Reactions Worksheet eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 20 Oxidation Reduction Reactions Worksheet Compatibility with Devices
  - Chapter 20 Oxidation Reduction Reactions Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 20 Oxidation Reduction Reactions Worksheet
  - Highlighting and Note-Taking Chapter 20 Oxidation Reduction Reactions Worksheet
  - Interactive Elements Chapter 20 Oxidation Reduction Reactions Worksheet
8. Staying Engaged with Chapter 20 Oxidation Reduction Reactions Worksheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 20 Oxidation Reduction Reactions Worksheet
9. Balancing eBooks and Physical Books Chapter 20 Oxidation Reduction Reactions Worksheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 20 Oxidation Reduction Reactions Worksheet
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 20 Oxidation Reduction Reactions Worksheet
  - Setting Reading Goals Chapter 20 Oxidation Reduction Reactions Worksheet
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 20 Oxidation Reduction Reactions Worksheet
  - Fact-Checking eBook Content of Chapter 20 Oxidation Reduction Reactions Worksheet
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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