

Chapter 1 – Kinetics - Practice Test #1

Name - _____

The answers to this test are posted online

1. Consider the following reaction mechanism:



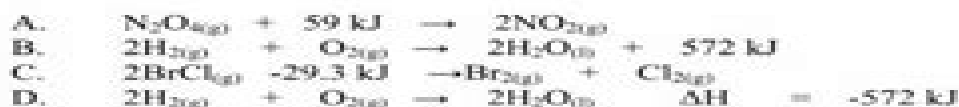
The chemical species MX is a(n)

- A. catalyst
B. inhibitor
C. final product
D. reaction intermediate
2. Consider the following reaction: $2\text{N}_2\text{O}_{5(g)} \rightarrow 4\text{NO}_{2(g)} + \text{O}_{2(g)}$
At a certain temperature the rate of decomposition of $2\text{N}_2\text{O}_{5(g)}$ is 2.5×10^{-6} mol/s. The rate of formation of NO_2 is
- A. 1.0×10^{-5} mol/s
B. 1.3×10^{-6} mol/s
C. 2.5×10^{-6} mol/s
D. 5.0×10^{-6} mol/s

3. Which of the following factors affect the rates of both homogeneous and heterogeneous reactions.

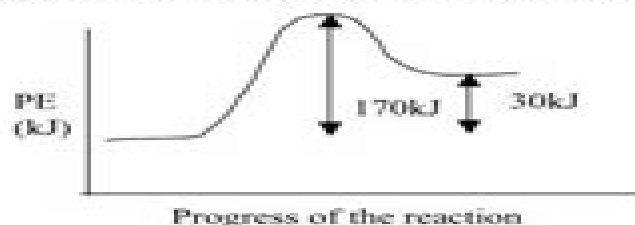
- I nature of the reactants
II presence of a catalyst
III temperature of system
IV concentration of reactants

- A. I and IV only
B. II and III only
C. II, III, and IV only
D. I, II, III, and IV
4. Which of the following equations represents an endothermic reaction?



5. Consider the potential energy diagram. The activation energy for the reverse reaction is

- A. 30 kJ
B. 140 kJ



- C. 170 kJ
D. 200 kJ

Chapter 1 chemical Kinetics Test

Walther Nernst



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