

## Chemistry Worksheets Class 10 on Chapter 1 Chemical Reactions and Equations - Set 3

Q-1: Write the equations for the following reactions in the ionic form and balance the ionic equations.

- i) Magnesium is treated with dilute sulphuric acid
- ii) Stannous chloride solution is added into mercuric chloride solution
- iii) H<sub>2</sub>S gas is passed into copper(II) sulphate solution.
- iv) Liquid bromine is added to the container of sodium iodide crystals
- v) Aluminium is treated with dilute hydrochloric acid

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- a) Aluminium reacts with NaOH to form and
- b) In the activity series of metal, aluminium is active than iron.
- c) Copper sulphate solution is \_\_\_\_\_ in nature.
- slows down the process of oxidation in packaged food.

Q-3: What is the correct representation of the reaction that occurs during the chlor alkali process?

- a)  $2NaCl(I) + 2H_2O(I) \rightarrow 2NaOH(I) + Cl_2(g) + H_2(g)$
- b)  $2NaCl(aq) + 2H_2O(aq) \rightarrow 2NaOH(aq) + Cl_2(g) + H_2(g)$
- c)  $2NaCl(aq) + 2H_2O(l) \rightarrow 2NaOH(aq) + Cl_2(aq) + H_2(aq)$
- d)  $2NaCl(aq) + 2H_2O(l) \rightarrow 2NaOH(aq) + Cl_2(g) + H_2(g)$

Q-4: Which of the following statements is correct?

- a) Metals act as reducing agent and non metals act as oxidising agent
- b) Metals act as oxidising agent and non metals acts as reducing agent
- c) Metals and nonmetals both acts as reducing agents
- d) Metals and nonmetals both acts as oxidising agents

Q-5: What do you expect to happen when the AgNO<sub>3</sub> solution is added to i) NaCl solution ii) CCl<sub>2</sub>? Give masons.

Q-6: Give reasons for the following

- Zinc and Aluminium are reactive metals and still they don't corrode easily.
- Bronze metals get coated with a greenish layer if not stored/cleaned properly.

Q-7: The following aqueous solutions are given to you:

ZnSO<sub>4</sub>, PbSO<sub>4</sub>, Ni(SO<sub>4</sub>)<sub>2</sub> and Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>

You are also given vessels made of Zn, Pb , Al and Ni.

# <u>Chapter 1chemical Reactions Practice Problems Answer</u> <u>Key</u>

Steven S. Zumdahl

## **Chapter 1chemical Reactions Practice Problems Answer Key:**

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