

KEY

Use the following to answer questions 1 & 2.



1. If 24 grams of sodium chloride reacts with an excess amount of magnesium oxide, how many grams of sodium oxide will be produced?

A	B
Given: 24 g NaCl = ? grams Na ₂ O	

Conversions: Molar mass of NaCl: 1 mole NaCl = 58 g NaCl
Mole to mole ratio: 2 mole NaCl = 1 mole Na₂O
Molar mass of Na₂O: 1 mole Na₂O = 62 g Na₂O

Setup:

$$24 \text{ g NaCl} \times \frac{1 \text{ mole NaCl}}{58 \text{ grams NaCl}} \times \frac{1 \text{ mole Na}_2\text{O}}{2 \text{ mole NaCl}} \times \frac{62 \text{ g Na}_2\text{O}}{1 \text{ mole Na}_2\text{O}} = 12.8 \text{ g Na}_2\text{O}$$

2. If 53 grams of magnesium oxide reacts with an excess amount of sodium chloride, how many grams of magnesium chloride will be made?

	A	B
Given:	53 g MgO	= ? g MgCl ₂

Conversions: Molar mass of MgO: 1 mole MgO = 40 g MgO
Mole to Mole Ratio : 1 mole MgO = 1 mole MgCl₂
Molar mass of MgCl₂ : 1 mole MgCl₂ = 94 g MgCl₂

Setup:

$$53 \text{ g MgO} \times \frac{1 \text{ mole MgO}}{40 \text{ g MgO}} \times \frac{1 \text{ mole MgCl}_2}{1 \text{ mole MgO}} \times \frac{94 \text{ g MgCl}_2}{1 \text{ mole MgCl}_2} = 124.55 \text{ g MgCl}_2$$

Use the following to answer questions 3 & 4.



3. If 14 moles of potassium oxide reacts with an excess amount of magnesium nitride, how many grams of potassium nitride will be made?

Given: 14 grams $\overset{\text{A}}{\text{K}_2\text{O}} = ? \overset{\text{B}}{\text{g K}_3\text{N}}$

Conversions: Molar mass of K_2O : 1 mole K_2O = 94 g K_2O
Mole to mole ratio: 3 moles K_2O = 2 moles K_3N
Molar mass K_3N : 1 mole K_3N = 131 g K_3N

Setup:

$$14 \text{ grams K}_2\text{O} \times \frac{1 \text{ mole K}_2\text{O}}{94 \text{ g K}_2\text{O}} \times \frac{2 \text{ moles K}_3\text{N}}{3 \text{ moles K}_2\text{O}} \times \frac{131 \text{ g K}_3\text{N}}{1 \text{ mole K}_3\text{N}} = 13 \text{ g K}_3\text{N}$$

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Stoichiometry Key

Léo Pomar



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