



Limiting Reactants and Theoretical Yield, *continued*

Sample Problem E Solution, *continued*

$$\begin{aligned} ? \text{ g H}_3\text{PO}_3 &= 225 \text{ g PCl}_3 \times \frac{1 \text{ mol PCl}_3}{137.32 \text{ g PCl}_3} \times \frac{1 \text{ mol H}_3\text{PO}_3}{1 \text{ mol PCl}_3} \\ &\quad \times \frac{82.00 \text{ g H}_3\text{PO}_3}{1 \text{ mol H}_3\text{PO}_3} = 134 \text{ g H}_3\text{PO}_3 \end{aligned}$$



Concept Review Continued Limiting Reactants And Percentage Yield

James E. Brady, John R. Holum



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