

Chapter 2 Atoms, Molecules and Ions

As you study this chapter, prepare an Atomic Theory Timeline with the experiment or idea that led to an advancement, the person or people associated with this, and the approximate date. We will add to the timeline in later chapters.

X To include: Democritus, Dalton, Thomson, Proust, Lavoisier, Becquerel, Millikan, Rutherford, Chadwick

Section 2.1 Atomic Theory

Dalton's Atomic Theory 1808

Three main ideas: Atoms, compounds, and chemical reactions

Proust - Law of Definite Proportions. Different samples of the same compound always contain the same elements in the same mass ratio.



All samples of carbon dioxide have the same ratio of mass of oxygen to mass of carbon.

123 g CO₂ sample:

89.4 g O / 33.6 g C = 2.66 g O : g C

or

88.6 g CO₂ sample:

64.4 g / 24.2 g C = 2.66 g O : g C

Law of multiple proportions: If two elements can combine to form more than one compound, the masses of one element that combine with a fixed mass of the other element are in ratios of small whole numbers.

Example 1: Xenon and fluorine can combine to form several different compounds, including XeF₂, which contains 0.2894 g F for every gram of Xe. Use the law of multiple proportions to determine *n*, which represents the number of F atoms in another compound, XeF_n, given that it contains 0.8982 g of F for every gram of Xe.

$$n = \frac{\text{g F: 1.000 g Xe}}{\text{g F: 1.000 g Xe}} = \frac{0.8982}{0.2894} \approx 3.104 \approx 3 \quad (\text{3 atoms of F in XeF}_n \text{ or XeF}_3)$$

So, the formula is: XeF₃

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Answer

Susan McMurry, John McMurry



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