

## CHAPTER 2- ACIDS AND BASES

### Questions And Answers 2023-2024

According to the Bronsted-Lowry definition, a base is a proton \_\_\_\_\_, while an acid is a proton \_\_\_\_\_. -  
Correct Answer-acceptor; donor

According to the Bronsted-Lowry definition, an acid must contain which of the following? - Correct Answer-  
A hydrogen atom

A Bronsted-Lowry base must have an "available" electron pair to form a bond to  $H^+$ . Which of the following structural features allow an electron pair to be used in this way? - Correct Answer-- One or more electron pairs in  $\pi$  bonds.

- One or more lone pairs in the structure.

Which of the following Bronsted-Lowry acid/base classifications are correct? - Correct Answer--  $(CH_3)_3N$  is a base.

-  $CH_3COOH$  is an acid.

Select all statements that accurately describe Bronsted-Lowry acid-base reactions. - Correct Answer--  
Acid-base reactions are also known as proton transfer reactions.

- One bond is formed in an acid-base reaction.

Match each term with the correct Bronsted-Lowry definition.

- Acid:

- Base: - Correct Answer-- proton ( $H^+$ ) donor

- proton ( $H^+$ ) acceptor

When identifying Bronsted-Lowry acids and bases in an organic reaction, if only one species has a lone pair or  $\pi$  bond it must be the \_\_\_\_\_, while if only one species contains an available hydrogen atom it must be the \_\_\_\_\_. - Correct Answer-base, acid

\*Reason: A species with a lone pair can readily accept a proton and is therefore defined as a base.

Which of the following statements correctly describe Bronsted-Lowry acids? - Correct Answer-- A Bronsted-Lowry acid must contain a H atom.

- A Bronsted-Lowry acid is a proton donor.

# Chapter 2acid And Bases Answers

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