

Focus 1

The properties of gases

Answers to exercises

Topic 1A The perfect gas

- 1A.1** (a) 810 Torr
 (b) 0.962 atm
 (c) 0.222 atm
 (d) 1.03×10^5 Pa
- 1A.2** 92 kPa
- 1A.3** 2.25 kPa
- 1A.4** 4.33 mmol
- 1A.5** 660 bar
- 1A.6** 10.0 atm
- 1A.7** 418 kPa
- 1A.8** 173 kPa
- 1A.9** -243.6°C
- 1A.10** 394 K
- 1A.11** (a) 3.6 m^3
 (b) 178 m^3
- 1A.12** $3.4 \times 10^5\text{ m}^3$
- 1A.13** benzene 0.355, toluene 0.645
- 1A.14** (a) $1.32 \times 10^{-3}\text{ m}^3$
 (b) $61.2 \times 10^3\text{ Pa}$

Topic 1B The kinetic model of gases

- 1B.1** (a) 493 m s^{-1}
 (b) 615 m s^{-1}
- 1B.2** (a) $v_{\text{mean}}(79\text{ K}) = 647\text{ m s}^{-1}$
 $v_{\text{mean}}(315\text{ K}) = 1.29\text{ km s}^{-1}$
 $v_{\text{mean}}(1500\text{ K}) = 2.82\text{ km s}^{-1}$
 (b) $v_{\text{mean}}(79\text{ K}) = 323\text{ m s}^{-1}$
 $v_{\text{mean}}(315\text{ K}) = 645\text{ m s}^{-1}$
 $v_{\text{mean}}(1500\text{ K}) = 1.41\text{ km s}^{-1}$
- 1B.3** 3.72
- 1B.4** m_{N_2} 0.80 g
 m_{He} 0.30 g
- 1B.5** 65 mPa

- 1B.6** $\lambda = 0.97\text{ }\mu\text{m}$
 $N = 120 \times 10^3$
- 1B.7** (a) 7.0×10^{10}
 (b) 7.0×10^8
 (c) 7.0×10^4
- 1B.9** (a) 6.8 nm
 (b) 68 nm
 (c) 6.8 mm

Topic 1C Real gases

- 1C.1** $z_c = 0.242$
- 1C.2** CH_4 $z_c = 0.288$
 C_2H_6 $z_c = 0.274$
 C_3H_8 $z_c = 0.276$
- 1C.3** 2.15 MPa
- 1C.4** $-3.37 \times 10^{-5}\text{ m}^3\text{ mol}^{-1}$
- 1C.5** $-3.49 \times 10^{-8}\text{ Pa}^{-1}$
- 1C.6** (a) (i) 101 kPa
 (ii) 83.1 MPa
 (b) (i) 101 kPa
 (ii) 206 MPa
- 1C.7** -0.608 MPa
- 1C.8** $b = 4.6 \times 10^{-5}\text{ m}^3\text{ mol}^{-1}$
 $Z = 0.66$
- 1C.9** 1842 K

Answers to problems

- 1.1** 0.50 m^3
- 1.2** 28.82 g mol^{-1}
 130 kg
- 1.3** 6.8 kPa
- 1.4** 713 Torr
- 1.5** $x_{\text{N}_2} = 0.7809$
 $x_{\text{O}_2} = 0.2094$
 $x_{\text{Ar}} = 0.0098$
- 1.6** 132 g mol^{-1}
- 1.7** 16.4 g mol^{-1}

Chapter 1section The Properties Of Gases Answers

**David Halliday,Robert Resnick,Jearl
Walker**



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