$660 \quad 7.22 \times 10^{10}$ 050 3.01 × 10<sup>24</sup> (c)  $1.20 \times 10^{10}$ 14 0.53 mol H 1.5 0.250 mol 16 (a) 262.87 g mol<sup>-1</sup> (b) 176.14 g moh! (c) 164.10 g mot-1 (d) 248.22 g mol-1 189.1.0 117 1.5-mol 1.6 0.0074 mol CF 1.83 × 10<sup>(4)</sup> C atoms 20 171 g (integer value because no calculator) 10.0 g HJO 2121 2.0 mol N<sub>2</sub> > 3.0 mol NH<sub>2</sub> > 25.0 mol H<sub>2</sub> > 1.0 mol N,H, (a) CH 650 CHLO (c) C\_H\_O... (d) C,H, (e) C.H. 00 CHLO Na,5,0,1 26 CoSO\_7HLO 27 C\_H\_N 28 SAR-EL 29 6.94 Lt Costs 3401 34 empirical formula CH; molecular formula C,H, empirical formula H.PO.; molecular formula H.P.O., C.,H.,N,P,O., for both empirical and molecular flormulass. CHLO Let y = mass of chalk in grams. moles of chalk used = mass used

M/(Ca/CO.)

These calculations have used  $L \approx 6.02 \times 10^{10}$ 

This is the same as the number of moles of carbon atoms used.

Therefore the number of carbon atoms used = moles of chalk × (6.02 × 1023 mol-1) =  $9.02 \times 10^{10} V$ 

- \$40 2.50 mol
  - 4ct 665.5 a.

100.09

- 37 (a) 2C,H, + 13O, → 8CO, + 10H,O (b) 1.59 g
- .4.355 kg
- (a) CaCO, -+ CaC + CO,
  - (b) 92.8%
  - (c) CaCO, is the only source of CO,; all the CaCO, undergoes complete decomposition; all CO, released is captured; heating does not cause any change in the mass of the other minerals present.
- 40 (a) 85.2 g
- (b) 1.3 g H.

(b) 5.63 mol

- 5.23 g CJH, CL
- 253.2 a theoretical CaSO.; 77.9%
- 3.16 g ester
- 107 g of C.H., needed
- 45 (a) 2.40 mol (c) 44 mol

Rorm O.E.F.O.O. 6d9

- (a) 35.65 dm\* (b) 5.7 dm<sup>2</sup>
- 0.652 dm<sup>3</sup>
- 0.138 mol Br., and 0.156 mol Cl., so more impleoules of Ct,
- 0.113 dm<sup>3</sup>
- 0.28 dm<sup>3</sup>
- 90 kPs
- 16.10 W-25
- 3.0 dm<sup>2</sup>
- 2.8 dm<sup>3</sup>

3

34.500 100,09 g mol\*\*

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**David Hendricks** 

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