

Ch 14.6-14.7 – Buffers & titrations
and
Ch 15 – Equilibria of Other Reaction Classes

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Section 14.6 – Buffers

What is a **buffer**? What two components are required to make a buffer?

Consider a buffer that contains 0.250 M HF and 0.200 M F^- .

What component in the buffer neutralizes any added base (OH^-)?

What component in the buffer neutralizes any added acid (H_3O^+)?

For buffer solutions, an ICE table may be used to determine equilibrium concentrations. Follow Example 14.20 (part a) to see how this is done. Create an ICE table for a buffer that is 0.250 M HF and 0.200 M F^- .

	HF (aq)	+	H ₂ O (l)		F ⁻ (aq)	+	H ₃ O ⁺ (aq)
I			--				
C			--				
E			--				

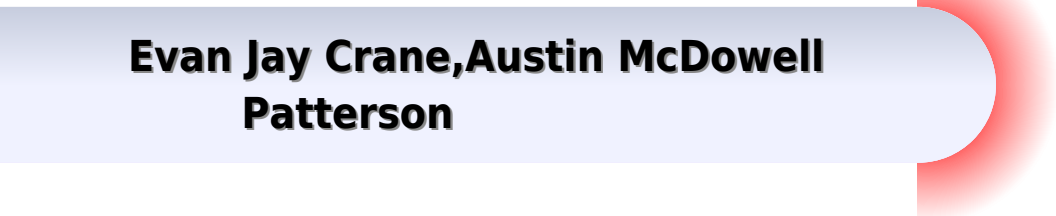
Q1: Use your ICE table to calculate the pH of this buffered solution. The K_a of HF is 3.5×10^{-4} . You may assume x is negligibly small.

The pH of **buffered solutions** can also be calculated using the Henderson-Hasselbalch equation. Write this equation below.

pH =

Chapter 1 stoichiometry Reading Guide

**Evan Jay Crane, Austin McDowell
Patterson**



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