

Nuclear Chemistry - Guided Practice Questions**Multiple Choice** - Choose the one best answer and write its letter in the blank provided.

_____ 1) In balancing the nuclear reaction: $^{238}_{92}\text{U} \rightarrow ^{234}_{90}\text{Th} + \text{_____}$, the identity of the missing element is _____?

- A) Pu B) U C) Th D) Np E) Pa

_____ 2) The following rxn, $^{210}_{84}\text{Po} \rightarrow ^{210}_{82}\text{Pb} + \text{_____}$, is an example of what form of radioactive decay?

- A) alpha B) positron C) gamma D) beta E) e-capture

_____ 3) The missing product from this reaction is: $^{138}_{54}\text{Xe} \rightarrow ^{138}_{52}\text{Te} + \text{_____}$?

- A) $^0_{-1}\text{e}$ B) $^0_0\gamma$ C) ^4_2He D) ^1_0n E) ^1_1p

_____ 4) What is the mass number of a neutron _____?

- A) 1 B) 3 C) 2 D) 0 E) 4

_____ 5) In the nuclear decay reaction represented by $^{238}_{92}\text{U} \rightarrow ^{234}_{90}\text{Th}$, the emitted particle is _____.

- A) beta particle B) neutron C) alpha particle D) positron E) proton

_____ 6) The reaction shown below is responsible for creating C-14 in the atmosphere. What is the bombarding particle _____?



- A) electron B) proton C) positron D) alpha particle E) neutron

_____ 7) $^{131}_{53}\text{I}$ has a half-life of 8.04 days. Assuming you start with a 1.53 mg sample of $^{131}_{53}\text{I}$, how many mg will remain after 13.0 days _____?

- A) 0.440 B) 0.422 C) 0.268 D) 0.835 E) 0.499

_____ 8) All atoms of a given element have the same _____.

- A) atomic mass B) mass number C) # of nucleons D) atomic # E) # of neutrons

_____ 9) Atoms containing radioactive nuclei are called _____.

- A) radioisotopes B) radionuclides C) nuclides D) nucleons E) radionophores

_____ 10) What happens to the mass number and the atomic number of an element when it undergoes beta decay?

- A) The mass number decreases by 4 and the atomic number decreases by 2.
 B) The mass number does not change and the atomic number decreases by 2.
 C) Neither the mass number nor the atomic number change.
 D) The mass number increases by 2 and the atomic number increases by 1.
 E) The mass number does not change and the atomic number increases by 1.

Chapter 28 Nuclear Chemistry Practice Problems Answers

Robert W. Avakian



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Table of Contents Chapter 28 Nuclear Chemistry Practice Problems Answers

1. Understanding the eBook Chapter 28 Nuclear Chemistry Practice Problems Answers
 - The Rise of Digital Reading Chapter 28 Nuclear Chemistry Practice Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 28 Nuclear Chemistry Practice Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 28 Nuclear Chemistry Practice Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 28 Nuclear Chemistry Practice Problems Answers
 - Personalized Recommendations
 - Chapter 28 Nuclear Chemistry Practice Problems Answers User Reviews and Ratings
 - Chapter 28 Nuclear Chemistry Practice Problems Answers and Bestseller Lists

5. Accessing Chapter 28 Nuclear Chemistry Practice Problems Answers Free and Paid eBooks
 - Chapter 28 Nuclear Chemistry Practice Problems Answers Public Domain eBooks
 - Chapter 28 Nuclear Chemistry Practice Problems Answers eBook Subscription Services
 - Chapter 28 Nuclear Chemistry Practice Problems Answers Budget-Friendly Options
6. Navigating Chapter 28 Nuclear Chemistry Practice Problems Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 28 Nuclear Chemistry Practice Problems Answers Compatibility with Devices
 - Chapter 28 Nuclear Chemistry Practice Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 28 Nuclear Chemistry Practice Problems Answers
 - Highlighting and Note-Taking Chapter 28 Nuclear Chemistry Practice Problems Answers
 - Interactive Elements Chapter 28 Nuclear Chemistry Practice Problems Answers
8. Staying Engaged with Chapter 28 Nuclear Chemistry Practice Problems Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 28 Nuclear Chemistry Practice Problems Answers
9. Balancing eBooks and Physical Books Chapter 28 Nuclear Chemistry Practice Problems Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 28 Nuclear Chemistry Practice Problems Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 28 Nuclear Chemistry Practice Problems Answers
 - Setting Reading Goals Chapter 28 Nuclear Chemistry Practice Problems Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 28 Nuclear Chemistry Practice Problems Answers
 - Fact-Checking eBook Content of Chapter 28 Nuclear Chemistry Practice Problems Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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