

9 Calculations from chemical equations

Answers to end-of-chapter questions

Pages 163–164 Questions

1. molar mass of $\text{Fe}_2(\text{SO}_4)_3 = [2 \times 55.8] + 3 \times [32.1 + (4 \times 16.0)] = 399.9 \text{ g mol}^{-1}$
molar mass of $\text{Fe}(\text{OH})_3 = 55.8 + (3 \times 17.0) = 106.8 \text{ g mol}^{-1}$
amount of $\text{Fe}_2(\text{SO}_4)_3$ (in moles) = mass/molar mass = $12.7 \text{ g} / 399.9 \text{ g mol}^{-1} = 0.03176 \text{ mol}$
moles $\text{Fe}(\text{OH})_3$ = moles $\text{Fe}_2(\text{SO}_4)_3 \times$ number of $\text{Fe}(\text{OH})_3$ in equation/number of $\text{Fe}_2(\text{SO}_4)_3$ in equation
= $0.03176 \text{ mol} \times 2 = 0.06352 \text{ mol}$
mass of $\text{Fe}(\text{OH})_3$ = moles \times molar mass = $0.06352 \text{ mol} \times 106.8 \text{ g mol}^{-1} = 6.78 \text{ g}$

[e] It is a good idea to work out the molar masses of the two substances involved before starting the rest of the calculation.

2. molar mass of $\text{AgNO}_3 = 107.9 + 14.0 + (3 \times 16.0) = 169.9 \text{ g mol}^{-1}$
amount (in moles) of AgNO_3 = mass/molar mass = $12.6 \text{ g} / 169.9 \text{ g mol}^{-1} = 0.07416 \text{ mol}$
ratio of $\text{Cu}:\text{AgNO}_3 = 1:2$
so, amount of copper = $\frac{1}{2} \times 0.07416 = 0.03708 \text{ mol}$
mass of copper needed = moles \times molar mass
= $0.03708 \text{ mol} \times 63.5 \text{ g mol}^{-1} = 2.35 \text{ g}$

[e] The answer 2.35 g was obtained keeping all the numbers on the calculator during the calculation. If the rounded up value of 0.0371 mol is used, the answer 2.36 g is obtained. Either would score full marks.

A common error is to calculate the molar mass of AgNO_3 as $2 \times 170 \text{ g mol}^{-1}$ because there are 2 moles of it in the equation. This is wrong. The number of moles of a substance depends only on its mass and not on the reaction. The stoichiometry comes into play when moles of one substance are converted to moles of another substance. This type of calculation can be done either by mass ratio or by converting to moles, then using the reaction stoichiometry and finally converting back to mass. The second method fits all types of calculation, so is the better one to use.

3. molar mass of $\text{NaNO}_3 = 23.0 + 14.0 + (3 \times 16.0) = 85.0 \text{ g mol}^{-1}$
amount of NaNO_3 = mass/molar mass = $33.3 \text{ g} / 85.0 \text{ g mol}^{-1} = 0.392 \text{ mol}$
ratio of $\text{O}_2:\text{NaNO}_3 = 1:2$
so, amount of oxygen = $\frac{1}{2} \times 0.392 = 0.196 \text{ mol}$
volume of oxygen = moles \times molar volume
= $0.196 \text{ mol} \times 25.0 \text{ dm}^3 \text{ mol}^{-1}$
= 4.90 dm^3
4. ratio of $\text{NH}_3:\text{H}_2 = 2:3$

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Christian Drosten



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1. Understanding the eBook Concept Review Measurements And Calculations In Chemistry Answer Key
 - The Rise of Digital Reading Concept Review Measurements And Calculations In Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Concept Review Measurements And Calculations In Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Concept Review Measurements And Calculations In Chemistry Answer Key
 - User-Friendly Interface

4. Exploring eBook Recommendations from Concept Review Measurements And Calculations In Chemistry Answer Key
 - Personalized Recommendations
 - Concept Review Measurements And Calculations In Chemistry Answer Key User Reviews and Ratings
 - Concept Review Measurements And Calculations In Chemistry Answer Key and Bestseller Lists
5. Accessing Concept Review Measurements And Calculations In Chemistry Answer Key Free and Paid eBooks
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 - Concept Review Measurements And Calculations In Chemistry Answer Key Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Concept Review Measurements And Calculations In Chemistry Answer Key Compatibility with Devices
 - Concept Review Measurements And Calculations In Chemistry Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Concept Review Measurements And Calculations In Chemistry Answer Key
 - Highlighting and Note-Taking Concept Review Measurements And Calculations In Chemistry Answer Key
 - Interactive Elements Concept Review Measurements And Calculations In Chemistry Answer Key
8. Staying Engaged with Concept Review Measurements And Calculations In Chemistry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Concept Review Measurements And Calculations In Chemistry Answer Key
9. Balancing eBooks and Physical Books Concept Review Measurements And Calculations In Chemistry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Concept Review Measurements And Calculations In Chemistry Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Concept Review Measurements And Calculations In Chemistry Answer Key
 - Setting Reading Goals Concept Review Measurements And Calculations In Chemistry Answer Key
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Concept Review Measurements And Calculations In Chemistry Answer Key
 - Fact-Checking eBook Content of Concept Review Measurements And Calculations In Chemistry Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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