

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

| | |
|---|-------------------------|
| _____ lattice energy _____ | a. ionic bond |
| _____ bond energy _____ | b. covalent bond |
| _____ enthalpy of vaporization _____ | c. metallic bond |

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

| | | | |
|-----------------------------------|-----------------|--------------------------------------|---------------|
| _____ ionic _____ | a. H—F | _____ nonpolar covalent _____ | d. H—H |
| _____ ionic _____ | b. Na—Cl | _____ polar covalent _____ | e. H—C |
| _____ polar covalent _____ | c. H—O | _____ polar covalent _____ | f. H—N |

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen
atom and the unshared electron pair of a strongly electronegative atom such as O,
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally
intermolecular, while ionic and covalent bonds occur between ions or atoms
respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H₂O bonds means a stronger intermolecular attraction, making
water a liquid at room temperature. Hydrogen bonding exists between water
molecules, but not between hydrogen sulfide molecules.

Covalent Bonding 8 1 Study Guide Answer Key

Russell C. Hollingsworth



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Table of Contents Covalent Bonding 8 1 Study Guide Answer Key

1. Understanding the eBook Covalent Bonding 8 1 Study Guide Answer Key
 - The Rise of Digital Reading Covalent Bonding 8 1 Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding 8 1 Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Covalent Bonding 8 1 Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding 8 1 Study Guide Answer Key
 - Personalized Recommendations
 - Covalent Bonding 8 1 Study Guide Answer Key User Reviews and Ratings

- Covalent Bonding 8 1 Study Guide Answer Key and Bestseller Lists
- 5. Accessing Covalent Bonding 8 1 Study Guide Answer Key Free and Paid eBooks
 - Covalent Bonding 8 1 Study Guide Answer Key Public Domain eBooks
 - Covalent Bonding 8 1 Study Guide Answer Key eBook Subscription Services
 - Covalent Bonding 8 1 Study Guide Answer Key Budget-Friendly Options
- 6. Navigating Covalent Bonding 8 1 Study Guide Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Covalent Bonding 8 1 Study Guide Answer Key Compatibility with Devices
 - Covalent Bonding 8 1 Study Guide Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonding 8 1 Study Guide Answer Key
 - Highlighting and Note-Taking Covalent Bonding 8 1 Study Guide Answer Key
 - Interactive Elements Covalent Bonding 8 1 Study Guide Answer Key
- 8. Staying Engaged with Covalent Bonding 8 1 Study Guide Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonding 8 1 Study Guide Answer Key
- 9. Balancing eBooks and Physical Books Covalent Bonding 8 1 Study Guide Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonding 8 1 Study Guide Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Bonding 8 1 Study Guide Answer Key
 - Setting Reading Goals Covalent Bonding 8 1 Study Guide Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Bonding 8 1 Study Guide Answer Key
 - Fact-Checking eBook Content of Covalent Bonding 8 1 Study Guide Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
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