

Class 11 Chemical Bonding and Molecular Structure Important Questions with Answers

Short Answer Type Questions

1. Explain the non linear shape of H_2S and non planar shape of PCl_3 using valence shell electron pair repulsion theory.

Answer.

- The main atom in H_2S is S, which has two lone pairs. These lone pairs cause repulsion and displace the H-S bond, resulting in a non-linear shape.



- PCl_3 has a trigonal planar structure. P has three single bonds and one lone pair (pair of unshared electrons). Each chlorine atom has a single 3p orbital that is completely occupied. The overlap of a phosphorus sp^3 hybrid orbital with a singly occupied chlorine 3p orbital results in the formation of P-Cl bonds. Three lone pairs are held by each Cl atom.



2. Using molecular orbital theory, compare the bond energy and magnetic character of O_2^+ and O_2^- species.

Answer.

The electronic configurations of O_2^+ and O_2^- species are as follows:

Covalent Bonding And Molecular Structure Lab Answers

William G. O'Neal



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