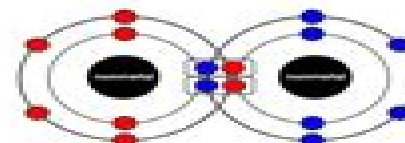


DRAWING COVALENT BONDS

Introduction:

A **covalent bond** forms between **nonmetals**. Covalent bonds involve **sharing pairs of electrons** between two atoms so that each atom gains a full outer valence shell. Sharing 1 pair of electrons creates a **single covalent bond**, sharing 2 pairs of electrons creates a **double covalent bond**, and sharing 3 pairs of electrons creates a **triple covalent bond**.



How to Draw Simple Covalent Bond Diagrams	
1.	Draw the electron dot diagrams for each element. <i>HINT: This will help you see the number of unpaired valence electrons for each element.</i>
2.	Count the total number of valence electrons between all atoms. <i>HINT: The total number of valence electrons should be 8 in most cases, except 2 hydrogens. If there are not 8, add more atoms of the element with the least number of valence electrons to start.</i>
3.	Determine which element is the central atom in the molecule, if needed. <i>HINT: The central atom is typically the atom with the largest valence number or the atom that there is only one of in the formula. If there are only 2 atoms in the formula, there is no central atom!</i>
4.	Pair up single electrons from different atoms to form covalent bonds and circle each pair. <i>HINT: Keep doing this until each single electron from all atoms have been paired up. You may need to use double or triple bonds to make this work. This step takes trial and error. If one arrangement is not working, try adding another element or trying a new structure. Practice makes perfect!</i>

Activity: Drawing Covalent Bond Diagrams:

- For each covalent compound, you have been given the element names and the chemical formula.
- Follow the steps listed above to create a covalent bonding diagram for the compound.
- Remember, this takes trial and error. The first 2 have been done for you as examples.

#	Nonmetal	Nonmetal	Formula	Bonding Diagram
A.	hydrogen $\text{H} \cdot$ <i>(must share 1)</i>	chlorine $\cdot \ddot{\text{Cl}} \cdot$ <i>(must share 1)</i>	HCl <i>(Formula shows 1 H and 1 Cl)</i>	
B.	hydrogen $\text{H} \cdot$ <i>(must share 1)</i>	oxygen $\cdot \ddot{\text{O}} \cdot$ <i>(must share 2)</i>	H_2O <i>(Formula shows 2 H and 1 O)</i>	

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discussing how they should and should not be described It reviews both theoretical developments and experimental procedures in order to explore interactions between nonbonded entities in such a fundamental manner as to elucidate their nature and origins Drawing attention to the extensive experience of its editor and team of expert authors *Intra and Intermolecular Interactions Between Noncovalently Bonded Species* is an indispensable guide to the foundational knowledge latest advances most pressing challenges and future directions for all those whose work is influenced by these interactions Comprehensively describes the nature of interactions between nonbonded species in biological systems liquids crystals clusters and in particular water Combines fundamental theoretical background information based on various approximations with the knowledge of experimental techniques Outlines interactions clearly and consistently with a particular focus on frequency and time resolved spectroscopies as applied to these interactions General College Chemistry Charles William

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