Chapter 9 Review

- 1) What does stoichiometry mean? "measuring elements"
- 2) Before performing a stoichiometry problem, you must first do what with the chemical equation? balance it
- 3) The coefficients in a chemical equation represent what?moles
- 4) In a chemical equation, what relationships are shown?molar ratios
- 5) Solve the following mole-mole conversions using this equation:

2NaHCO₂ | Na₂CO₃ + CO₂ + H₂O

a) If 3.5 moles of sodium bicarbonate are decomposed, how many moles of carbon dioxide are produced?

b) If 1.29 moles of water are produced by the decomposition of sodium bicarbonate, how many moles of sodium carbonate are produced?

c) How many moles of sodium bicarbonate are required to produce 19.5 moles of carbon dioxide?

- Use the equation in # 5 to solve the following mass-mass conversions.
 - a) If given 10.5 grams of sodium bicarbonate, how many grams of sodium carbonate can be produced?

b) How many grams of sodium bicarbonate would be required to produce 9.0 grams of water?

7) What conversion factor (mole ratio) would you use to solve for moles of sulfur formed from sulfur dioxide in the reaction

$$SO_2 \rightarrow S + O_2$$

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