

Unit 7 Study Guide:

Atoms, Elements/Compounds/Mixtures, States of Matter, Physical and Chemical Properties and Changes

1. Famous scientists in the study of the atom: Name, dates, and contribution. See summary on page 203.

	Date	Contribution
Aristotle	425BC	Earth, Air, Fire, Water, Aether
Democritus	425BC	Matter is made of indivisible atoms
John Dalton	1802	Atomic Theory of Matter
Joseph Priestly	1772	Discovered Oxygen
Antoine Lavoisier	1783	1 st Chemistry Textbook, L.O.C.M.
J.J. Thomson	1897	Discovered the electron
Ernest Rutherford	1909	Discovered the proton
Niels Bohr	1913	Proposed electron energy levels
Gilbert Lewis	1916	Proposed the term 'valence'
James Chadwick	1932	Discovered the neutron

2. Using the Periodic Table to find the characteristics of an atom:

Consider Phosphorous P-15. See practice pages 211, 213, 218.

	P	Bohr Model	Lewis Model
Atomic Symbol:	P		
Atomic Number:	15		
Atomic Mass:	30.97		
Protons:	15		
Electrons:	15		
Neutrons:	16		
Group:		5	
Period:	2		
Electron Shells:	3		
Valence Electrons:	5		
Ionic State:	-3		
Type:	non metal		
Most common Isotope:	P-31		



3. Classifying matter; mixture or pure substance, homogeneous or heterogeneous, and compound or element.

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A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the text 'Open University. S334 Course Team'.

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