

## CHAPTER 21 REVIEW

**Nuclear Chemistry****SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ Based on the information about the three elementary particles in the text, which has the greatest mass?  
(a) the proton  
(b) the neutron  
(c) the electron  
(d) They all have the same mass.
2. \_\_\_\_\_ The force that keeps nucleons together is  
(a) a strong nuclear force.  
(b) a weak nuclear force.  
(c) an electromagnetic force.  
(d) a gravitational force.
3. \_\_\_\_\_ The stability of a nucleus is most affected by the  
(a) number of neutrons.  
(b) number of protons.  
(c) number of electrons.  
(d) ratio of neutrons to protons.
4. \_\_\_\_\_ If an atom should form from its constituent particles,  
(a) matter is lost and energy is taken in.  
(b) matter is lost and energy is released.  
(c) matter is gained and energy is taken in.  
(d) matter is gained and energy is released.
5. \_\_\_\_\_ For atoms of a given mass number, those with greater mass defects, have  
(a) smaller binding energies per nucleon.  
(b) greater binding energies per nucleon.  
(c) the same binding energies per nucleon as those with smaller mass defects.  
(d) variable binding energies per nucleon.
6. Based on Figure 5.3 of the text, which isotope of He, helium-3 or helium-4,  
\_\_\_\_\_ a. has the smaller binding energy per nucleon?  
\_\_\_\_\_ b. is more stable to nuclear changes?
7. The number of neutrons in an atom of magnesium-25 is \_\_\_\_\_.
8. Nuclides of the same element have the same \_\_\_\_\_.

# Chemistry Ch 21 Answers Study Guide

**Marc Loudon, Jim Parise, Lisa Bonner**



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