

26/05/2016

Q.P. Code : 28592

(2. Hours)

[Total Marks : 60

- N.B. : (1) Question No. 1 is compulsory.
 (2) Answer any three questions from remaining five questions.
 (3) Figure to the right indicates full marks.
 (4) Atomic weights : Ca = 40, Mg = 24, Cl = 35.5, S = 32, H = 1, O = 16.

1. Attempt any five from the following : 15
 - (a) Write two balanced equations to describe the changes that occur when hard water is boiled.
 - (b) Give reasons to explain why natural rubber needs vulcanization.
 - (c) Give the preparation and uses of silica bricks.
 - (d) Give the number of phases in the following system (any three) :
 - (i) Saturated solution of NaCl
 - (ii) Mixture of rhombic and monoclinic sulphur
 - (iii) Mixture of O_2 and N_2
 - (iv) $Ice \rightleftharpoons Water$ equilibrium
 - (e) What is grease ? What are the conditions in which greases are used?
 - (f) Thermosetting polymers cannot be reshaped and reused. Give reasons.
 - (g) Calculate the COD of an effluent sample if 25c.c. of the effluent sample required 8.3 c.c. of 0.001M $K_2Cr_2O_7$ for oxidation.
2. (a) Calculate the quantity of lime and soda required for softening 50,000 L of water containing following salts per litre. 6

$Ca(HCO_3)_2 = 16.2mg$; $Mg(HCO_3)_2 = 7.5mg$;
 $CaSO_4 = 13.6mg$; $MgSO_4 = 24.0mg$.
 $MgCl_2 = 10.0mg$.
- (b) Explain the following terms : 5
 - (i) Condensed Phase rule
 - (ii) Triple point
- What are carbon nanotubes ? Describe the laser method of preparation of CNT. 4

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