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## Writing Balanced Equations CHEM 10 Review Worksheet

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The equations supplied here are fairly simple and could be considered "Chem 10 level". They are provided to assist your review of some topics covered in Chp 3 of the Zumdahl textbook. An answer key is provided on the second page.

### Writing Balanced Chemical Equations

Write balanced chemical equations for each of the following reactions:

1. Gaseous nitrogen dioxide reacts with water to produce nitric acid and gaseous nitrogen monoxide.
2. Rubidium metal (*s*) + water (*l*) → rubidium hydroxide (*aq*) + hydrogen (*g*)
3. Iodine (*s*) + chlorine (*g*) → iodine trichloride (*l*)
4. Sodium bicarbonate when heated will form solid sodium carbonate, carbon dioxide and water vapor.
5. Magnesium acetate (*aq*) + ammonium sulfide (*aq*) → magnesium sulfide (*s*) + ammonium acetate (*aq*).
6.  $\text{C}_6\text{H}_{14}$  (*l*) + oxygen (*g*) → carbon dioxide (*g*) + water (*l*)      [ $\text{C}_6\text{H}_{14}$  is named hexane]
7. Iron(III) oxide (*s*) + carbon (*s*) → iron (*s*) + carbon dioxide (*g*)
8. Potassium carbonate (*aq*) + nitric acid (*aq*) → potassium nitrate (*aq*) + carbon dioxide (*g*) + water (*l*)
9. Ammonium nitrite (*s*) → nitrogen (*g*) + water (*l*)
10. Gaseous sulfur dioxide reacts with oxygen forming gaseous sulfur trioxide.
11. Cobalt(III) bromide (*aq*) reacts with pure chlorine forming pure bromine and cobalt(III) chloride (*aq*).
12. Silver (*s*) +  $\text{S}_8$  (*s*) → silver sulfide (*s*)      [ $\text{S}_8$  is a naturally occurring form of sulfur]
13. Ammonia (*g*) + oxygen (*g*) → nitrogen monoxide (*g*) + water (*l*)
14. Aqueous barium hydroxide reacts with phosphoric acid forming water and solid barium phosphate.
15. Aluminum (*s*) + copper(II) sulfate (*aq*) → aluminum sulfate (*aq*) + copper (*s*)
16. When heated, solid lithium chlorate decomposes to produce solid lithium chloride and oxygen gas.
17. Benzene liquid ( $\text{C}_6\text{H}_6$ ) burns in oxygen to form carbon dioxide and water (and heat).
18. Silicon dioxide (*s*) + hydrofluoric acid (*aq*) → silicon tetrafluoride (*g*) + water (*l*)

# Concept Review Writing Balanced Equations Answers

**Lingsheng Yao**



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