

Study Card: Solvation

What Does Dissolving Involve?

When a solid dissolves in water, it's not *gone*. It has changed its state. As a solid, a solute particle (ion, molecule, etc.) is attracted to other solute particles. But when it dissolves in H_2O , it changes to the aqueous state and becomes attracted to the H_2O molecules.

The Dissolving of Ionic Solids

Dissolving involves more than just *mixing*. When table salt (NaCl) is added to water, the ions break off the crystal one-by-one and are surrounded by H_2O molecules. The O atom of H_2O surrounds the + ions. The H atoms of H_2O surround the - ions. This is known as **solvation**; it prevents the ions from returning to the crystal lattice. All ionic solids are dissolved in this manner.

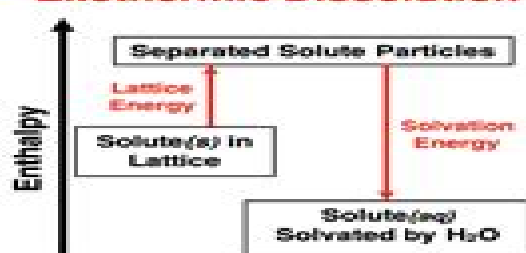
The Dissolving of Molecular Solids

Molecular solids consist of molecules instead of ions. When table sugar (sucrose, $\text{C}_{12}\text{H}_{22}\text{O}_{11}$) dissolves, the molecules break off the crystal one-by-one and are surrounded by H_2O molecules. Hydrogen bonding forces hold the sucrose molecules together. So H_2O surrounds the -OH groups of a sucrose molecule, forming its own hydrogen bonds with sucrose and preventing the molecule from returning to the crystal. This is how **solvation** helps molecular solids dissolve.

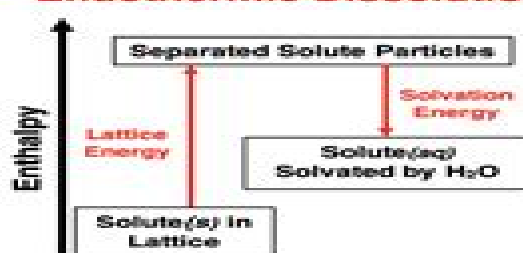
An Energy Perspective on Dissolving

Dissolving can be thought of as involving two steps. First, solute particles are separated from the crystal. This requires energy; the amount is roughly equal to the **lattice energy**. Second, the H_2O molecules surround the solute particles. This releases energy; the amount is roughly equal to the **solvation energy**. Dissolution is more likely to occur if the solvation energy is greater than the lattice energy.

Exothermic Dissolution



Endothermic Dissolution



The Dissolving Process



Solute particles
attract each other.

Solute particles
attract water.

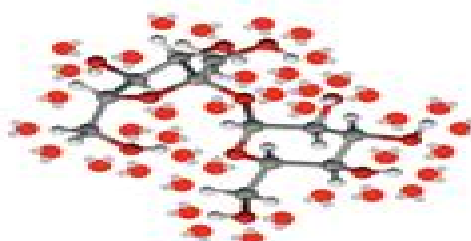
Solvation

Polar H_2O molecules surrounding + and - ions.



Solvation

H_2O molecules attract and surround sucrose.



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Richard H. Moyer



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