

Conceptual Chemistry

PRACTICE PAGE

Chapter 10: Acids and Bases In Our Environment

Expert-Only Questions ♦♦♦

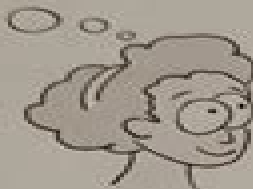
1. To say that an acid is strong in aqueous solution means that the acid
- produces a large concentration of hydronium ions.
 - is quite odorous when dissolved in water.
 - does not form chemical bonds with water.
 - does not separate into ions when dissolved in water.

$$10^3 = 1000$$

$$10^2 = 100$$

$$10^1 = 10$$

$$10^0 = ?$$



2. What does it mean to say that a solution has a pH of zero?
- There are no hydronium ions in that solution.
 - The concentration of hydronium ions is 1 M.
 - The solution is very alkaline.
 - It is not possible to have a pH equal to zero!

3. We measure the "acidity" of a solution by measuring the concentration of hydronium ions, H_3O^+ . What does it mean to say that one has a "neutral" solution, such as one has with pure water?

- There are no hydronium ions in this solution.
- The concentration of hydronium ions is very small.
- The concentration of hydronium ions equals the concentration of hydroxide ions, OH^- .



Pure?



4. Water reacting with itself is an endothermic reaction:



What happens, therefore, to the pH of neutral water as its temperature is increased and thus more thermal energy becomes available? The pH of neutral water

- remains at 7 because the pH of neutral water is always 7.
- decreases to some value below 7.
- increases to some value above 7.

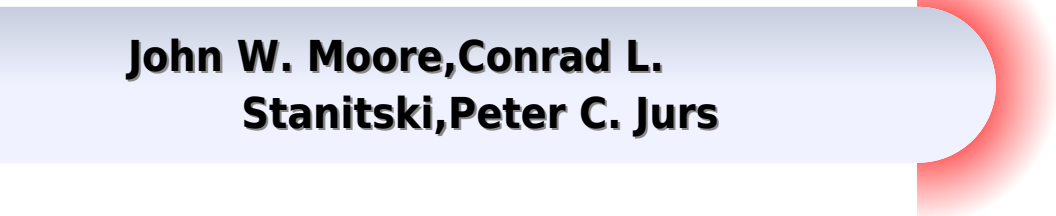
The reaction of water with itself is favored by greater temperatures and pressures. At the super-high temperature and pressure of 374°C and 218 atmospheres, the reaction of water with water is so great that K_w equals 10^{-13} , which is ten times greater than the usual value of 10^{-14} . What is the pH of such a solution? Is this pH acidic, neutral, or basic? What might you predict about the corrosive properties of water under these conditions? So much to learn about our universe!



S. Subak

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**John W. Moore, Conrad L.
Stanitski, Peter C. Jurs**



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