

In this activity we will learn how to draw Lewis structures, calculate formal charge and examine resonance. We will also observe the Octet rule and exceptions to this rule.

#### AP Chem. Lewis Dot Structure Worksheet

Write Lewis Dot Structures and calculate the formal charges for the following compounds:

- 1. SO<sub>2</sub>
- 2. CO<sub>3</sub>-2
- 3. NH<sub>2</sub>
- 4. CH<sub>4</sub>
- 5. H<sub>2</sub>
- 6. HF
- 7. H<sub>2</sub>O
- 8. N<sub>2</sub>
- 9. HC1
- $10.H_2S$
- 11.PF<sub>3</sub>
- 12.CH3COOH
- 13.NH<sub>4</sub>\*
- 14.NO<sub>2</sub>°
- $15.SO_4^{-2}$
- 16.NO<sub>2</sub>\*

- 18. H<sub>2</sub>O<sub>2</sub>
- 19. C<sub>2</sub>H<sub>6</sub>
- 20. CH<sub>3</sub>OH
- 21. C<sub>2</sub>H<sub>4</sub>
- 22. C<sub>4</sub>H<sub>10</sub>
- 23. C2H4OH
- 24. CIO
- 25. H<sub>2</sub>S<sub>2</sub>
- 26. HNO<sub>3</sub>
- 27. H<sub>2</sub>SO<sub>4</sub>
- 28. CH<sub>3</sub>NO<sub>2</sub>
- 29. C<sub>2</sub>H<sub>2</sub>
- 30. CIO<sub>2</sub>:
- 31. PO<sub>4</sub>-3
- 32. CO
- 33. CO<sub>2</sub>

17.OH

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