Chemistry - Unit 6 Notes 1 - Thomson Model of the Atom

J. J. Thomson performed experiments with cathode rays in an attempt to understand electricity – which was still a mystery in the late 1800s. Use the website A Look Inside the Atom! to find the conclusions that Thomson and other physicists drew regarding the mysterious cathode rays.

<u>Thomson's 1897 Experiments</u> - state the conclusions Thomson drew from each of his famous cathode ray experiments:

- First Experiment: Thomson directed the beam at an electrometer and tried to separate
 the evidence of charge from the path of the beam. What connection did Thomson find
 between charge and the cathode rays? Was the charge positive or negative?
- 2. Second Experiment: Thomson tried passing the cathode ray through an electric field. How did cathode ray beam behave when it passed through an electric field? What did he conclude after his second experiment?
- 3. Third Experiment: Thomson did some careful measurements on how much the path of the cathode ray was bent in a magnetic field and how much energy they carried. From this work Thomson could describe the mass/charge ratio of the cathode ray particles. What amazing result did Thomson find?

Thomson's Atomic Model: Thomson presented three hypotheses from his experiments. Only two were accepted by physicists — in fact the third was shown to be wrong! From the first two came a model of the atom known as the *Plum Pudding* model. Complete the atom drawing below to illustrate Thomson's plum pudding model. Explain how this fits with his observations.



https://www.aip.org/history/exhibits/electron/jjrays.htm @Modeling Instruction - AMTA 2013

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Lab Manual for General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, 2009-08-21 Teaching all of the necessary concepts within the constraints of a one term chemistry course can be challenging Authors Denise Guinn and Rebecca Brewer have drawn on their 14 years of experience with the one term course to write a textbook that incorporates biochemistry and organic chemistry throughout each chapter emphasizes cases related to allied health and provides students with the practical quantitative skills they will need in their professional lives Essentials of General Organic and Biochemistry captures student interest from day one with a focus on attention getting applications relevant to health care professionals and as much pertinent chemistry as is reasonably possible in a one term course Students value their experience with chemistry getting a true sense of just how relevant it is to their chosen profession. To browse a sample chapter view sample ChemCasts and more visit www whfreeman com gob Manual Nitrogen Control ,1993 Item no 0431 K in Education ,1984-10 Active Learning Patricia Ann Mabrouk, 2007 This symposium series book focuses on the application of active learning methods in teaching analytical science broadly defined at both the undergraduate and graduate levels The volume includes a wide range of examples of how these methods are being applied at public and private community colleges four year colleges and graduate research universities in the United States and abroad Solutions Manual to Accompany Physical Chemistry J. Philip Bromberg, 1984 Study Guide and Solutions Manual for Organic Chemistry Susan McMurry, 1992 John McMurry s best selling text presents organic chemistry in a new edition that is up to date beautifully written visually striking and pedagogically sound Described by many of its users as an eminently teachable text McMurry sets the standard in the field The writing style has received almost universal acclaim from its users McMurry introduces new concepts only as needed and immediately illustrates them with concrete examples And wherever possible he ties material together with brief reviews overviews and reaction summaries. The result is a text that helps students mentally organize the material a text that helps them understand concepts not just memorize facts and a text that helps them make sense of the voluminous amount of material they encounter in the study of organic chemistry McMurry uses a simple but important polar reaction the addition of HBr to an alkene as the lead off reaction to illustrate the general principles of organic reactions Users of former editions found this an excellent choice because of its relative simplicity no prior knowledge of chirality or kinetics is required and its importance as a polar reaction on a common functional group that offers students the key to understanding hundreds of thousands of ionic reactions By selecting this particular model McMurry is able to offer The International Military Digest Annual ,1918 an unusually early presentation of organic reactions Catalog Issue of the Maryville College Bulletin Maryville College, 1919 Annual Report of the Normal, Model, Grammar, and Common Schools in Upper Canada Ontario. Dept. of Education, 1921 Votes & Proceedings New South Wales. Parliament. Legislative Council, 1904 Journal of the Legislative Council New South Wales. Parliament. Legislative Council, 1904

Annual Report of the President of the University on Behalf of the Regents to His Excellency the Governor of the State of California University of California, Berkeley, 1902 **Interim Report of the Commissioners on Certain Parts of** Primary Education New South Wales. Commission on primary, secondary, technical, and other branches of education, 1903 Register - University of California University of California, Berkeley, 1914 Register University of California, Berkelev, 1916 Register ... California. University, University of California, Berkeley, 1914 Circular of Information with Reference Primarily to the Undergraduate Division at Berkeley University of California, Berkeley, 1916 Catalogue of the College of California and College School University of California (System),1914 **Study Guide and Solutions** Manual to Accompany Fundamentals of Organic Chemistry McMurry, Susan McMurry, 1990 Building Age and National Builder ,1928

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