

Stoichiometry Test

Name: _____

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Learning (Product) 12 18 7 11

Assessment of

Multiple Choice: Select the most correct answer. (5X)

1. Which is the correct mole ratio of K_3PO_4 to HNO_3 in the unbalanced chemical reaction?



- A) 1:1 B) 2:3 C) 1:9 D) 1:2

2. The limiting reactant of a chemical reaction is

- A) the reactant of a chemical process that is not consumed completely
B) the product of a chemical process that is consumed completely
C) the product of a chemical process that is not consumed completely
D) the reactant of a chemical process that is consumed completely
E) none of the above

3. Potassium chloride and oxygen gas are formed upon heating potassium chlorate. The following balanced equation shows this decomposition process:



The number of moles of $KClO_3$ required to produce 5.0 mol of oxygen is

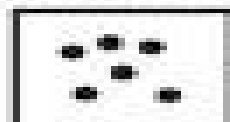
- A) 1.0 B) 2.0 C) 3.0 D) 3.3 E) 5.0

4. Consider the following unbalanced reaction: $Ca + H_2O \rightarrow H_2 + CaOH$

Which statement about the above reaction is true?

- A) 2 g of Ca will react with 2 moles of water
B) 2 moles of calcium will produce 1 moles of hydrogen gas
C) 2 moles of water will produce 1 g of hydrogen
D) 2 g of water will produce 1 g of hydrogen

5. Which is true of the reaction shown below?



+



→



● = Substance X

○ = Substance Y

●○ = Substance Z

A) The mole ratio of this reaction is 6:5:6.

Chemistry Test On Stoichiometry With Answers

**Peter J. Mikulecky, Michelle Rose
Gilman, Kate Brutlag**



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