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## Chemistry - Balancing Equations

Balance the following equations by inserting the proper coefficients. For selected reactions, draw *Before* and *After* particle diagrams to show the particles involved in the reaction. Be sure to provide a key.

18.3

188

- 1 C + 1 H<sub>2</sub>O → 1 CO + 1 H<sub>2</sub>
- 2. 2 MgO  $\rightarrow$  2 Mg + 1 O2
- 3. 4 Al + 3 O<sub>2</sub> → 2 Al<sub>2</sub>O<sub>3</sub>



- 4. 1 Zn + 1  $H_2SO_4 \rightarrow 1 ZnSO_4 + 1 H_2$
- 1 Cl<sub>2</sub> + 2 KI → 2 KCl + 1 I<sub>2</sub>
- 2 CuCl → 2 Cu + 1 Cl2
- 2 Na + 1 Cl<sub>2</sub> → 2 NaCl
- 8. 2 Al + 6 HCl → 2 AlCl<sub>3</sub> + 3 H<sub>2</sub>
   • • • • • • •



- 9.  $2 \text{ Fe}_2\text{O}_3 \rightarrow 4 \text{ Fe} + 3 \text{ O}_2$
- 10. 4 P + 5 O2 → 2 P2O5
- 11.  $\underline{1}$  Mg +  $\underline{2}$  HCl  $\rightarrow$   $\underline{1}$  MgCl<sub>2</sub> +  $\underline{1}$  H<sub>2</sub>
- 12.  $3 \text{ H}_2 + 1 \text{ N}_2 \rightarrow 2 \text{ NH}_3$
- 13. 1 BaCl<sub>2</sub> + 1 H<sub>2</sub>SO<sub>4</sub>  $\rightarrow$  1 BaSO<sub>4</sub> + 2 HCl
- 14. 1 CH<sub>4</sub> + 2 O<sub>2</sub> → 1 CO<sub>2</sub> + 2 H<sub>2</sub>O

  ••• ••• •••



#14 Before After

# **Chemistry Unit 7 Reaction Equations 1 Answers**

John S. Phillips, Cheryl Wistrom

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