

Chemistry C2 1.1-2.5

Atomic Structure

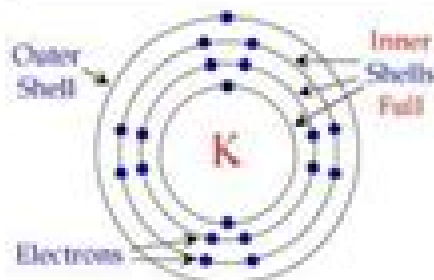
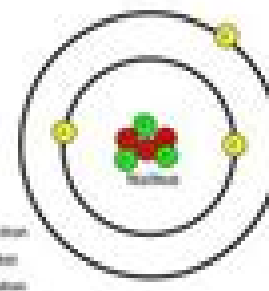
The element **Lithium** has a symbol of **Li**.

3 electrons (atomic number) and an atomic mass of 7

3
Li
6.94

Atoms are made of **protons**, **neutrons** and **electrons**.

Protons have a **positive +** charge and electrons have a **negative -** charge. Atoms are arranged in the periodic table in order of their atomic number. Neutrons have **no** electric charge so are **neutral**.



The electrons in an atom are arranged in **shells**.

Atoms with the same number of electrons in their outer shell belong in the same group of the periodic table. The number of electrons in the outer shell of an atom determines the way that the atom behaves in chemical reactions.

Chemical Bonding

Elements react to form compounds by **gaining** or **losing** electrons or by **sharing** electrons.

The elements in Group 1 react with the elements in Group 7 because group 1 needs to **lose** an electron from its outer shell and group 7 need to **gain** an electron to have a full outer shell.

This can be shown by using **dot and cross diagrams**.

Ionic Bonding

Ionic compounds are held together by strong forces between the oppositely charged ions. This is called **ionic bonding**.

Group 2 & Group 6 elements can form ionic compounds. This happens between a metal and non-metal. These form giant structures of ions in a lattice.

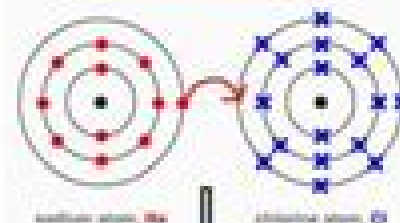
Covalent Bonding

This takes place between **non-metals**. Covalent bonds are formed when atoms **share** electrons. Many substances contain covalent bonds consist of molecules, but some have giant covalent structures. The electrons are shared so there is an outer shell of 8 electrons.

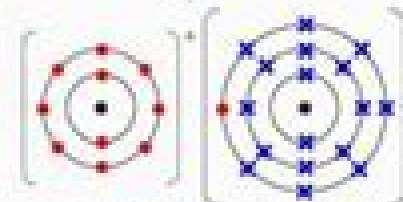
Remember: An outer shell which does not have a full shell of 8 electrons is **unstable**. So all elements are looking to become **stable**.



ionic bonding in sodium chloride



sodium atom, Na chlorine atom, Cl



sodium ion, Na⁺ chloride ion, Cl⁻

free electrons from outer shells of metal atoms



metal ions

Metallic Bonding

The atoms (or ions) in metals are arranged in regular layers. The positive ions in metals are held together by electrons from the outer shell of each metal atom. These form **giant** structures. The diagram on the left shows this.

Ionic compounds - it takes a lot of **energy** to break the bonds which hold a giant ionic lattice together. Very **high melting points** and are solid at room temperatures. They conduct electricity when melt or dissolve them in water because their ions can then move freely.

Chemthink Answers Of Atomic Structures

Jack Challoner



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1. Understanding the eBook Chemthink Answers Of Atomic Structures
 - The Rise of Digital Reading Chemthink Answers Of Atomic Structures
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemthink Answers Of Atomic Structures
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemthink Answers Of Atomic Structures
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemthink Answers Of Atomic Structures
 - Personalized Recommendations
 - Chemthink Answers Of Atomic Structures User Reviews and Ratings
 - Chemthink Answers Of Atomic Structures and Bestseller Lists
5. Accessing Chemthink Answers Of Atomic Structures Free and Paid eBooks
 - Chemthink Answers Of Atomic Structures Public Domain eBooks
 - Chemthink Answers Of Atomic Structures eBook Subscription Services
 - Chemthink Answers Of Atomic Structures Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Chemthink Answers Of Atomic Structures Compatibility with Devices
 - Chemthink Answers Of Atomic Structures Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemthink Answers Of Atomic Structures
 - Highlighting and Note-Taking Chemthink Answers Of Atomic Structures
 - Interactive Elements Chemthink Answers Of Atomic Structures
8. Staying Engaged with Chemthink Answers Of Atomic Structures
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemthink Answers Of Atomic Structures
9. Balancing eBooks and Physical Books Chemthink Answers Of Atomic Structures
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemthink Answers Of Atomic Structures
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemthink Answers Of Atomic Structures
 - Setting Reading Goals Chemthink Answers Of Atomic Structures
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemthink Answers Of Atomic Structures
 - Fact-Checking eBook Content of Chemthink Answers Of Atomic Structures
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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