

# ChemThink: Behavior of Gases

Name \_\_\_\_\_ Per \_\_\_\_\_

1. What does the speed of atoms depend on? Describe the relationship between speed and each of these factors.

temperature:  $\uparrow$  temp =  $\uparrow$  speed

type of atom: smaller/lighter = faster

2. In a sample of one type of atom, do all the atoms move at the same speed? Why?

No, speed is related to each individual atoms energy

3. What is the equation for Pressure and what are the units for pressure?

$$\text{Pressure} = \frac{\text{Force}}{\text{Area}} \quad \text{psi}$$

4. When atoms are moving in a closed container what is creating the force on the sides of the container?

impacts (collisions) of atoms w/ sides of container

5. Explain how a change in temperature causes pressure to change and include an equation.

more collisions results in higher pressure

$\downarrow$   
b/c molecules moving faster

$$\frac{P}{T} = \text{constant}$$

6. Draw the graph you observed of pressure vs. temperature. Is this a direct relationship or an inverse relationship?



direct

7. Fill the cause/effect relationship in the diagram of boxes you observed.



- Explain how the number of atoms present causes pressure to change and include an equation?

more atoms = more collisions, which increases pressure

$$\frac{P}{n} = \text{constant}$$

# Chemthink Behavior Of Gases Answers

**Brendan G. Carr**



### **Chemthink Behavior Of Gases Answers:**

*Chemistry and Our Universe* , In the first of two lectures on the properties of gases review the basic equations that describe their behavior Learn the history of Boyle s law Gay Lussac s law Charles s law and Avogadro s law Then use these four expressions to derive the celebrated ideal gas law      **Nonideal Behavior of Gases and Their Mixtures** Abdus Satter,1963

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