

Concept Review

Section: Limiting Reactants and Percentage Yield

Complete each statement below by choosing a term from the following list. Some may be used more than once.

excess limiting limiting actual theoretical
percentage excess theoretical

1. A(n) excess reactant is not completely used up in a chemical reaction.
2. A(n) limiting reactant is used up and thus controls the amount of theoretical that can be obtained in a chemical reaction.
3. The reactant that has not been used is the excess reactant.
4. The limiting reactant determines how much theoretical product can be obtained. The maximum amount of product is called the theoretical.
5. Limit is a factor in selecting the limiting reactant.
6. In a chemical reaction, percentage yield is usually less than 100% because theoretical is the maximum amount of product that can be obtained. In this case, the most important factor is theoretical yield vs. actual yield of the chemical reaction in this case.
7. The percentage yield is a way to measure reaction efficiency.
8. The percentage yield calculation tells you that percentage yield is the actual yield.
9. The percentage yield is the actual yield divided by the theoretical yield.
10. The percentage yield calculation tells you that percentage yield is the actual yield.

Answer the following items in the space provided.

11. When 100 g of A reacts with 100 g of B, 150 g of product is formed. What is the limiting reactant? What is the theoretical yield of the product?



Concept Review Answer Key Limiting Reactants And Percentage Yield

Arshad Iqbal



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Chemistry: Core Concepts, 3rd Edition Allan Blackman, Daniel Southam, Gwendolyn Lawrie, Natalie

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Prentice Hall Chemistry, 2000 *Improving Student Comprehension of Stoichiometric Concepts* Connie Lynn Bannick Kemner, 2007
General Chemistry, Reactions First Kevin Revell, 2024-12-04 Revell's General Chemistry empowers students to grasp essential topics and concepts with more ease Using a friendly approach the text uses metaphors and relatable examples to demystify even the most challenging subjects in general chemistry

Chemistry William R. Robinson, Jerome D. Odom, Henry Fuller Holtzclaw, 1992
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