The solubility rules allow us to determine if a doublereplacement reaction will actually occur.

Consider:

2 NaCl
$$(aq)$$
 + Mg(NO₃)₂ (aq) \rightarrow 2 NaNO₃ (aq) + MgCl₂ (aq)

Did a reaction actually occur? (Hint: Think about the 5 indicators of chemical change)

No. there is no reaction because all of the ions remain in solution.

- 1. Energy released
- 2. Gas evolution
- 3. Color change
- 4. Temperature change
- 5. Precipitate formation

How about:

$$2 \text{ NaCl } (aq) + \text{Pb(NO}_3)_2 (aq) \rightarrow 2 \text{NaNO}_3 (aq) + \text{PbCl}_2 (s)$$

Yes, because a precipitate was formed.

Double Replacement And Solubility Pogil Answers

Patrick Vollmar

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