

7. GROUPS 1 & 2

SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

- 1 Magnesium oxide, known as magnesia, can be produced by the thermal decomposition of magnesium carbonate.

The equation for the reaction is



- *(a) Other Group 2 carbonates also undergo thermal decomposition.

State and explain the trend in thermal stability of carbonates down Group 2.

(4)

Thermal stability increases down the group

Because ionic radius increases as you go down the group

This results in less polarisation (distortion) of the carbonate ion

And less weakening of the covalent bonds (C-O) within the carbonate ion

- *(c) State why carrying out a flame test is not an appropriate way to confirm the presence of magnesium ions in the sample of magnesia.

(1)

Magnesium ions give no flame colour

OR

Energy is emitted with Mg ions is outside of the visible region

(Total for Question 1 = 15 marks)

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Table of Contents Edexcel Biology 5bi1h Past Papers

1. Understanding the eBook Edexcel Biology 5bi1h Past Papers
 - The Rise of Digital Reading Edexcel Biology 5bi1h Past Papers
 - Advantages of eBooks Over Traditional Books
2. Identifying Edexcel Biology 5bi1h Past Papers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Edexcel Biology 5bi1h Past Papers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Edexcel Biology 5bi1h Past Papers
 - Personalized Recommendations
 - Edexcel Biology 5bi1h Past Papers User Reviews and Ratings
 - Edexcel Biology 5bi1h Past Papers and Bestseller Lists

5. Accessing Edexcel Biology 5bi1h Past Papers Free and Paid eBooks
 - Edexcel Biology 5bi1h Past Papers Public Domain eBooks
 - Edexcel Biology 5bi1h Past Papers eBook Subscription Services
 - Edexcel Biology 5bi1h Past Papers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Edexcel Biology 5bi1h Past Papers Compatibility with Devices
 - Edexcel Biology 5bi1h Past Papers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Edexcel Biology 5bi1h Past Papers
 - Highlighting and Note-Taking Edexcel Biology 5bi1h Past Papers
 - Interactive Elements Edexcel Biology 5bi1h Past Papers
8. Staying Engaged with Edexcel Biology 5bi1h Past Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Edexcel Biology 5bi1h Past Papers
9. Balancing eBooks and Physical Books Edexcel Biology 5bi1h Past Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Edexcel Biology 5bi1h Past Papers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Edexcel Biology 5bi1h Past Papers
 - Setting Reading Goals Edexcel Biology 5bi1h Past Papers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Edexcel Biology 5bi1h Past Papers
 - Fact-Checking eBook Content of Edexcel Biology 5bi1h Past Papers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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