

Exercises

2. Following the rule you established in the last key question, write correct chemical formula for each of the following compounds

Compound	Formula
(a) Lithium chloride	$\text{Li}^{+1} / \text{Cl}^{-1} \rightarrow \text{LiCl}$
(b) Magnesium iodide	$\text{Mg}^{+2} / \text{I}^{-1} \rightarrow \text{MgI}_2$
(c) Strontium selenide	$\text{Sr}^{+2} / \text{Se}^{-2} \rightarrow \text{SrSe}$
(d) Rubidium fluoride	$\text{Rb}^{+1} / \text{F}^{-1} \rightarrow \text{RbF}$
(e) Lithium oxide	$\text{Li}^{+1} / \text{O}^{-2} \rightarrow \text{Li}_2\text{O}$
(f) Sodium sulfide	$\text{Na}^{+1} / \text{S}^{-2} \rightarrow \text{Na}_2\text{S}$
(g) Potassium chloride	$\text{K}^{+1} / \text{Cl}^{-1} \rightarrow \text{KCl}$
(h) Calcium phosphide	$\text{Ca}^{+2} / \text{P}^{-3} \rightarrow \text{Ca}_3\text{P}_2$
(i) Barium oxide	$\text{Ba}^{+2} / \text{O}^{-2} \rightarrow \text{BaO}$
(j) Aluminum sulfide	$\text{Al}^{+3} / \text{S}^{-2} \rightarrow \text{Al}_2\text{S}_3$

3. Use your answers to the Key Questions and the Exercise Questions to draw a conclusion about the ratio of ions in two compounds if the elements in the compounds are from the same groups (example: aluminum oxide and aluminum sulfide; lithium chloride and potassium chloride).

Bonding Packet Answers

Robb H. Tracy

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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Table of Contents Bonding Packet Answers

1. Understanding the eBook Bonding Packet Answers
 - The Rise of Digital Reading Bonding Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Bonding Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Bonding Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Bonding Packet Answers

- Personalized Recommendations
- Bonding Packet Answers User Reviews and Ratings
- Bonding Packet Answers and Bestseller Lists
- 5. Accessing Bonding Packet Answers Free and Paid eBooks
 - Bonding Packet Answers Public Domain eBooks
 - Bonding Packet Answers eBook Subscription Services
 - Bonding Packet Answers Budget-Friendly Options
- 6. Navigating Bonding Packet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Bonding Packet Answers Compatibility with Devices
 - Bonding Packet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Bonding Packet Answers
 - Highlighting and Note-Taking Bonding Packet Answers
 - Interactive Elements Bonding Packet Answers
- 8. Staying Engaged with Bonding Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Bonding Packet Answers
- 9. Balancing eBooks and Physical Books Bonding Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Bonding Packet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Bonding Packet Answers
 - Setting Reading Goals Bonding Packet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Bonding Packet Answers

- Fact-Checking eBook Content of Bonding Packet Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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