Crude oil and biofuels	
(a) Gases, petrol, kerosene, diesel oil, fuel oil and bitumen are the fractions obtained from crude oil by fractional distillation.	
entify the fraction described in each of the following statements.	
This fraction is more difficult to ignite than most other fractions and is used as a fuel in large ships.	(1)
name of fraction	
This fraction is obtained from the top of the fractionating column.	(1)
name of fraction	
This fraction has a higher boiling point than kerosene and is used as a fuel for some cars.	(1)
name of fraction	
nen hydrocarbon fuels are burnt, several different products can be formed.	
nich of these cannot be a product of burning hydrocarbon fuels?	
t a cross (X) in the box next to your answer.	
	(1)
A carbon	
B carbon dioxide	
C hydrogen	
D water	
rbon monoxide gas, CO, can be formed when methane, CH _a , undergoes complete combustion.	
Balance this equation for the incomplete combustion of methane by putting numbers in the spaces provided.	(2)
$2CH_4 + \underline{\hspace{1cm}} O_2 \rightarrow 2CO + \underline{\hspace{1cm}} H_2O$	(2)
Carbon monoxide is a toxic gas and can cause death.	
Explain how carbon monoxide can cause death.	
	(2)
ofuels are produced from plants.	
	ses, petrol, kerosene, diesel oil, fuel oil and bitumen are the fractions obtained micrude oil by fractional distillation. Intify the fraction described in each of the following statements. This fraction is more difficult to ignite than most other fractions and is used as a fuel in large ships. name of fraction This fraction is obtained from the top of the fractionating column. name of fraction This fraction has a higher boiling point than kerosene and is used as a fuel for some cars. name of fraction the hydrocarbon fuels are burnt, several different products can be formed. In this fraction be a product of burning hydrocarbon fuels? It a cross (図) in the box next to your answer. A carbon B carbon dioxide C hydrogen D water Thom monoxide gas, CO, can be formed when methane, CH _{er} undergoes complete combustion. Balance this equation for the incomplete combustion of methane by putting numbers in the spaces provided. $2CH_4 + O_2 \rightarrow 2CO + O_2 + O_3 +$

(Total for Question 3 = 10 marks)

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