

Stoichiometry Worksheet #1

Mole-to-Mole Problems



How many moles of potassium nitrate are produced when 7.70 moles of aluminum nitrate react with excess potassium phosphate?

$\frac{7.70 \text{ mol } Al(NO_3)_3}{1 \text{ mol } Al(NO_3)_3}$	$\frac{3 \text{ mol } KNO_3}{1 \text{ mol } Al(NO_3)_3}$	=	$23.1 \text{ mol } KNO_3$
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How many moles of hydrogen are needed to completely react with 2.60 moles of nitrogen?

$\frac{2.60 \text{ mol } N_2}{1 \text{ mol } N_2}$	$\frac{3 \text{ mol } H_2}{1 \text{ mol } N_2}$	=	$7.80 \text{ mol } H_2$
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How many moles of ammonia are produced with 11.30 moles of nitrogen?

$\frac{11.30 \text{ mol } N_2}{1 \text{ mol } N_2}$	$\frac{2 \text{ mol } NH_3}{1 \text{ mol } N_2}$	=	$22.6 \text{ mol } NH_3$
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How many moles of oxygen are necessary to react completely with 4.0 moles of propane?

$\frac{4.0 \text{ mol } C_3H_8}{1 \text{ mol } C_3H_8}$	$\frac{5 \text{ mol } O_2}{1 \text{ mol } C_3H_8}$	=	$20 \text{ mol } O_2$
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Mole-to-Mass and Mass-to-Mole Problems



How many grams of barium chloride are needed to react completely with 5.00 moles of silver nitrate?

$\frac{5.00 \text{ mol } AgNO_3}{2 \text{ mol } AgNO_3}$	$\frac{1 \text{ mol } BaCl_2}{2 \text{ mol } AgNO_3}$	$\frac{208.23 \text{ g } BaCl_2}{1 \text{ mol } BaCl_2}$	=	$521 \text{ g } BaCl_2$
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Edgenuity Stoichiometry Answers

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