

- *(iii) Select a suitable indicator for this titration, giving the colour change you would expect to see.

Justify your selection.

(3)

- (c) The titration above can be used to determine the concentration of ammonia solution when ammonia is distributed between the two immiscible solvents, trichloromethane and water.

An experiment is carried out to find the equilibrium constant for the reaction:



75 cm³ of 4.0 mol dm⁻³ aqueous ammonia solution and 75 cm³ of the trichloromethane are shaken together. The two liquids are allowed to separate and 25.0 cm³ of the aqueous layer is taken and titrated with 4.0 mol dm⁻³ hydrochloric acid. The whole procedure is repeated.

The average titre is 24.0 cm³.

- (i) Calculate the number of moles of ammonia, and hence the concentration of ammonia, in mol dm⁻³, in the aqueous layer.

(3)



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G Thomas

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'G Thomas' text.

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