

Cumulative Exam Review

Consider each of the four major classes of elements. Give a strategy for identifying each class (with corresponding elements) and a strategy for identifying ions and typical Lewis acids. Which reactions are reversible?

- 1) Acid-base reactions: acids react with bases to form a conjugate base and a conjugate acid.
- 2) Oxidation-reduction reactions: oxidation is the loss of electrons, reduction is the gain of electrons.
- 3) Precipitation reactions: two aqueous solutions react to form an insoluble solid.
- 4) Complexation reactions: a metal ion reacts with a ligand to form a complex ion.

Identify the type of reaction for the following:

- 1) $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- 2) $\text{Fe} + \text{Cu}^{2+} \rightarrow \text{Fe}^{2+} + \text{Cu}$
- 3) $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$

Which of the following is a Lewis acid? Which is a Lewis base?

Which of the following is a Lewis acid? Which is a Lewis base?

- 1) H^+
- 2) OH^-
- 3) NH_3
- 4) BF_3

Which of the following is a Lewis acid? Which is a Lewis base?

- 1) H_2O
- 2) H_2O_2
- 3) H_2SO_4
- 4) H_2SO_3
- 5) H_2SO_2

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- 3) NH_3
- 4) BF_3

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