#### Double Replacement Reactions Practice 1.0 Answer New

Objective - - Classify chemical reactions and predict the products

- Apply the Law of Conservation of Matter to balance chemical equations

Directions: For each reaction below-

- A. Write the chemical equation, predicting the products as nameded.
- B. Balance the chemical equation.
- E. AgNOs + Kupos +

KaPO+ + 3 AgNO+ + 3 KNO+ + AgaPO+

2. Calcium acetate + melium carbonate

EarColloGyle + NasCOs + CaCOs + 2 NaCollyOs

3. Alaninum lodids v mercury(II) nitrate

Z Alba + 3 Hg(NO<sub>clic</sub> + 2 Al(NO<sub>clin</sub> + 3 Hgb)

4. Potassium chromate + magnesium chiorate

NiCros + MajCloub + MajCros + 2 8/20s

S.: Assessorium carbonate + tron(310) perchlorate

3 (NOS<sub>4</sub>)<sub>2</sub>CO<sub>6</sub> + 2 Pe(CEO<sub>4</sub>)<sub>5</sub> + Pe<sub>2</sub>(CO<sub>5</sub>)<sub>5</sub> + 6 NH<sub>4</sub>CIO<sub>4</sub>

6. Bajono - Nin -

Stadiotic - 2 Note - State - 2 Notottle

T. Sodium chlorate + Lead(32) acetate

2 Nath + Philliphiology - 2 Natiothio, + Phillip

B. Ammonium silicute + vanadium(V) sullice

5 (NHJ) (NO) + V(SO) + V(SO) + V(SO) + S (NHJ) (NO)

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Holt Rinehart & Winston, Holt, Rinehart and Winston Staff

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