

Ques. ► 6. $A = (n-1)d^5ns^1$, $B = (n-1)d^1ns^2$, Here $n = 4$

[D.B.-16]

- What is common-ion effect?
- In which orbital between 3d and 4p electron will enter first?
- Determine the four quantum number of the last electron of the element A.
- Which among the elements A and B will produce colored compound? Explain with reason.

Answer to the question no. 6

[a] The process by which the dissociation rate of the weak electrolyte is reduced due to the presence of a strong electrolyte is known as similar-ion effect.

[b] In which orbital the electron will enter first is determined by Aufbau's principle or $(n + l)$. According to this principle, the orbital of which $(n + l)$ value is low has low energy and electron will enter to it first. But if the $(n + l)$ value of two orbitals are same, then electron will enter the orbital whose n value is lower. For both 3d and 4p orbital the $(n + l)$ value is 5. But for 3d orbital the n value is 3 and for 4p orbital the n value is 4. So, electron will enter 3d orbital first.

[c] The last electron of the given element A enters into $(n-1)d$ orbital.

As, the value of n is 4.

So, the value of four quantum numbers of the last electron are :

(i) Principal quantum number $= n - 1 = 4 - 1 = 3$

(ii) The value of subsidiary quantum number, l for d orbital is 2.

(iii) Magnetic quantum number, $m = 2$.

Ques. ► 7



[D.B.-16]

- Write down the law of mass action.
- Which between H_2SO_3 and HNO_3 is more acidic and why?
- If the MN solution of the stem is saturated then calculate the solubility product of MN at $20^\circ C$.
- If the two solutions in the stem are mixed, will there be any ppt? Explain with reason.

Answer to the question no. 7

[a] At a fixed temperature the rate of any reaction is proportional to the active mass (partial pressure in case of liquid and concentration in case of gas) of its reactants. This is the law of mass action.

[b] Between H_2SO_3 and HNO_3 , HNO_3 is more acidic. Because we know that the acid whose oxidation number of the central atom is more is more acidic.

Here for H_2SO_3 , the oxidation number of the central atom S is

$+4$ and for HNO_3 , the oxidation number of the central atom N is $+5$. So, HNO_3 is more acidic than H_2SO_3 .

[c] The solubility product of MN at $20^\circ C$ is 3.5×10^{-4} . The concentration of MN in the solution is $0.30 M$. So, the solution is supersaturated and a precipitate will form.

[d] If the two solutions are mixed, the concentration of M^{2+} ions will be $0.15 M$ and the concentration of N^{2-} ions will be $0.005 M$. The ion product will be $0.15 \times 0.005 = 0.00075$, which is greater than the K_{sp} of PN (2.5×10^{-2}). So, a precipitate of PN will form.

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 - The Rise of Digital Reading Dowload Hsc Chemistry Paper 1notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Dowload Hsc Chemistry Paper 1notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Dowload Hsc Chemistry Paper 1notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Dowload Hsc Chemistry Paper 1notes
 - Personalized Recommendations
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 - Highlighting and Note-Taking Dowload Hsc Chemistry Paper 1notes
 - Interactive Elements Dowload Hsc Chemistry Paper 1notes
8. Staying Engaged with Dowload Hsc Chemistry Paper 1notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Dowload Hsc Chemistry Paper 1notes
9. Balancing eBooks and Physical Books Dowload Hsc Chemistry Paper 1notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Dowload Hsc Chemistry Paper 1notes
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Dowload Hsc Chemistry Paper 1notes
 - Setting Reading Goals Dowload Hsc Chemistry Paper 1notes
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Dowload Hsc Chemistry Paper 1notes
 - Fact-Checking eBook Content of Dowload Hsc Chemistry Paper 1notes
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
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 - Integration of Multimedia Elements
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