Name:

 Hydrogen, H<sub>2</sub>, is used as rocket fuel. The hydrogen is burned in oxygen, O<sub>2</sub>, according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H2 is burned?

2  $H_2(g) + O_2(g) \rightarrow 2 H_2O(g)$ ;  $\Delta H = -484 \text{ kJ}$ 

200.03 12 1 mol HZ = 99.2 mol Hz 99.2 mol Hz 1-484 KJ (2011)

Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide

 $CS_2(l) + 3O_2(g) \rightarrow CO_2(g) + 2SO_2(g) : \Delta H = -1077 \text{ kJ}$ What is  $\Delta H$  for the following reaction?

change 50 5:5 Not 1/2 CO2(8) + SO2(8) - 1/2 CS2(1) + 3/2 O2(8)

+ 1077 K3 = (538.5 K3)

3. Following is the equation for the combustion of acetone:

2 C<sub>3</sub>H<sub>6</sub>O(g) + 8 O<sub>2</sub>(g) - 6 H<sub>2</sub>O(g) + 6 CO<sub>2</sub>(g) + 853.6 kcal/2

How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

- 853. 6 Kc-1 2 mol Catte a

4.5 mol (3400 = (-1920.6 KC-1)

 The specific heat of iron is 0.110 cal / g.°C. How much heat must be added to 5.2 g of iron at 25 °C to raise its temperature to 225 °C? heat = 8 = m e AT = (5.2) (0.110 cml) (225-25'C) = (14.4)

 Hydrazine, N2H4, is a colorless liquid used as a rocket fuel. What is the enthalpy change for the Grant process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate  $\Delta H$  for reaction (1) given the data below.

> (1)  $N_2(g) + 2H_2(g) \rightarrow$

 $N_2H_4(1) + O_2(g) \rightarrow N_2(g) + 2H_2O(1); \Delta H = -622.2 \text{ kJ}$ 

 $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(f) : \Delta H = -285.8 \text{ kJ}$ 

N291+ 2420 (1) --> N2Hy (1) + O2(1, DH=622.2)

2H2G1 + OZG1 -> ZHZOQ1 DH= (Z)(-285.8) KJ NZG1 + ZHZG1 -> NZHYG1 (DH= + 50.6 KJ

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