

# CHEMISTRY



## Periodic Table of Elements

1 H 1.008																	18 Ar 39.948
2 He																	36 Kr
3 Li 6.941	4 Be 9.012															54 Xe	86 Rn
5 Na 22.990	6 Mg 24.305															84 Po	118 Og
7 K 39.098	8 Ca 40.078	9 Sc	10 Ti	11 V	12 Cr	13 Mn	14 Fe	15 Co	16 Ni	17 Cu	18 Zn	19 Ga	20 Ge	21 As	22 Se	23 Br	24 Kr
19 Rb	20 Sr	21 Y	22 Zr	23 Nb	24 Mo	25 Tc	26 Ru	27 Rh	28 Pd	29 Ag	30 Cd	31 In	32 Sn	33 Sb	34 Te	35 I	36 Xe
37 Cs	38 Ba	39 La	40 Ce	41 Pr	42 Nd	43 Pm	44 Sm	45 Eu	46 Gd	47 Tb	48 Dy	49 Ho	50 Er	51 Tm	52 Yb	53 Lu	54 Xe
55 Fr	56 Ra	57 Ac	58 Th	59 Pa	60 U	61 Np	62 Pu	63 Am	64 Cm	65 Bk	66 Cf	67 Es	68 Fm	69 Md	70 No	71 Lr	72 Hf
73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn	87 Fr	88 Ra	89 Ac	90 Th
91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr	104 Rf	105 Db	106 Sg	107 Bh	108 Hs
109 Mt	110 Ds	111 Rg	112 Cn	113 Nh	114 Fl	115 Mc	116 Lv	117 Ts	118 Og	119 Uut	120 Uuq	121 Uub	122 Uuq	123 Uub	124 Uuq	125 Uub	126 Uuq

1	ATOMIC NUMBER
H	SYMBOL
1.008	ATOMIC WEIGHT

- Elemental forms:**
  - Monoatomic:** Noble gases: helium, neon, argon, krypton, xenon, and radon; nonreactive
  - Diatoms:** Gases: fluorine, chlorine, hydrogen, oxygen, and nitrogen; Liquid: bromine; Solid: iodine
  - Molecular solids:** Sulfur
  - Covalent solids:** Carbon (diamond, graphite)
  - Metals:** Most common type of element; high coordination numbers; very dense
- Types of compounds:** Large variety due to ionic, covalent, and metallic bonding options
  - Ionic:** Electron exchange
  - Covalent:** Shared electrons
    - Polar covalent: Unequal sharing
  - Ion complex:** Central ion + Solvent
  - Transition metal compounds:**
    - Coordination compounds: Non-bonding ligands
    - Organometallics: Metal-C bond
  - Complex minerals:** Metal oxides and silicates
- Periodicity:** The chemical properties of an element depend on valence electrons
  - Family of elements:** Column in periodic table
  - The addition of d-orbitals produces an expanded octet

### Periodic Trends

- Chemistry determined by the behavior of valence electrons
- Atoms or ions get larger as you go down the column; electron shells are added
- Cation and anions (same charge) decrease in size as the atomic number ( $Z$ ) increases

- Ionization energies:** Noble gases are the most stable; alkali metals are the easiest to ionize; electron screening lessens  $Z$ -effect on orbital energy
- Atomic and ionic radii:** Not real properties; derive from structural data; divide bond into radii using a consistent scheme
  - Ionic radius:** Depends on charge, counter ion, and coordination (crystal type)
  - Atomic radius:** Based on diatomic gas data or metallic solid structural data

### Development of the Periodic Table

- Missing elements were predicted using periodicity
- Columns 1: Alkali metals
- Columns 2: Alkaline earth metals
- First transition series: Titanium-Copper
- Column 17: Halogens: fluorine, chlorine, bromine, iodine, and astatine
- Column 18: Rare gases: helium, neon, argon, krypton, xenon, and radon
- Artificial:  $Z = 90$  or above; radioactive
- Lanthanides: f-subshell elements; starts with  $Z = 58$



### Practical Application

- The colors of fireworks depend on the elements in the explosive

d-subshell 5 orbitals



## Atomic Structure

- Atomic number ( $Z$ ):** The number of protons in the nucleus; for a neutral atom,  $Z$  = number of electrons; for an anion of charge  $-q$ ,  $Z = +q$  electrons; for a cation of charge  $+q$ ,  $Z = -q$  electrons
- Mass number ( $A$ ):**  $A = Z + N$ , where  $N$  = the number of neutrons in the nucleus
- Isotopes:** Atoms with the same  $Z$ , different  $A$ ; an element can have one or more isotopes; any sample contains a number of isotopes; in practical work, use the average mass number or atomic weight (weighted average of natural isotopes for a given element)

First Ionization Potential vs.  $Z$



Atomic Radius vs.  $Z$



## Atomic Quantum Numbers & Orbitals

- $n$ : principle**
  - angular momentum (orbital shape)
  - quantum number (orbital direction)
  - For each  $n$ , the possible  $l$  values are  $0, 1, \dots, n-1$
  - For each  $l$ , the possible  $m_l$  values are  $-l, \dots, 0, \dots, +l$
  - $m_s$  has two possible values:  $+\frac{1}{2}$  and  $-\frac{1}{2}$  (spin up, spin down)
  - Each value of  $n$  denotes a "shell" in the atomic structure; each  $l$  denotes a subshell
- s-orbital:**  $l = 0$ , 1 type
- d-orbital:**  $l = 2$ , 5 types
- p-orbital:**  $l = 1$ , 3 types
- f-orbital:**  $l = 3$ , 7 types

### Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe all many-electron atoms and ions; a more rigorous treatment determines exact energy levels and orbital properties; the main result is that subshells have different energies; the filling of levels is guided by the Aufbau principle:

Aufbau Order for Filling Subshells



$n$	$l$	$m_l$	$m_s$	element
1	0	0	$+\frac{1}{2}$	H
1	0	0	$-\frac{1}{2}$	He
2	0	0	$+\frac{1}{2}$	Li
2	0	0	$-\frac{1}{2}$	Be
2	1	-1	$+\frac{1}{2}$	B
2	1	0	$+\frac{1}{2}$	C
2	1	1	$+\frac{1}{2}$	N
2	1	-1	$-\frac{1}{2}$	O
2	1	0	$-\frac{1}{2}$	F
2	1	1	$-\frac{1}{2}$	Ne

- Electron configuration:** Atomic orbital occupancy for an atom
- Pauli exclusion principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons: one with spin up,  $m_s = +\frac{1}{2}$ , and one with spin down,  $m_s = -\frac{1}{2}$
- Hund's rule:** Electrons fill p, d, and f subshells in a manner to maximize the number of unpaired electrons (half-filled orbitals); spin is maximized
- Ionization potential (IP):** Energy required to remove electron from an atom or ion
  - First IP:** Removal of first valence electron
- Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale, 0 for rare gas

# College Intro Chemistry Study Guide

**T. R. Dickson**



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Introductory Chemistry for Today Spencer L. Seager, Michael R. Slabaugh, 2004. Distinguished by its superior allied health focus and integration of technology, Seager and Slabaugh's *INTRODUCTORY CHEMISTRY FOR TODAY* Fifth Edition continues to lead the market on both fronts through numerous allied health related applications, examples, boxes, and a new Companion Web Site, *GOB ChemistryNow*™. In addition to the many resources found in *GOB ChemistryNow*, this powerful new Web site contains questions modeled after the Nursing School and Allied Health Entrance Exams and NCLEX LPN Certification Exams. The authors strive to dispel users' inherent fear of chemistry and to instill an appreciation for the role chemistry plays in our daily lives through a rich pedagogical structure and an accessible writing style that provides lucid explanations. In addition, Seager and Slabaugh's *CHEMISTRY FOR TODAY* Fifth Edition provides greater support in both problem solving and critical thinking skills. By demonstrating how this information will be important to a reader's future career and providing important career information online, the authors not only help readers to set goals but also to focus on achieving them.

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