

Colligative Properties, Homework 13

Table 1: Constants

Name	Formula	$k_b \frac{^{\circ}\text{C}}{m_{\text{particles}}}$	$k_f \frac{^{\circ}\text{C}}{m_{\text{particles}}}$
Water	H_2O	0.512	1.86
Naphthalene	C_{10}H_8	*****	6.9
Ethyl alcohol	$\text{C}_2\text{H}_5\text{O}$	1.22	1.99
Acetic acid	$\text{HC}_2\text{H}_3\text{O}_2$	3.07	3.90
Benzene	C_6H_6	2.53	4.90
Nitrobenzene	$\text{C}_6\text{H}_5\text{NO}_2$	5.24	7.00

- How many particles are generated when each of the following compounds are dissolved in water. a) FeCl_2 , b) NaCl , c) $\text{C}_6\text{H}_{12}\text{O}_6$ (glucose), d) $\text{Cr}_2(\text{SO}_4)_3$, e) CH_3NH_2 (methyl amine)
- Calculate the a) freezing point and b) boiling point when 78.75 g NaCl are dissolved in 275 g of water.
- Antifreeze in cars is usually made by mixing ethylene glycol, $\text{HOCH}_2\text{CH}_2\text{OH}$, with water to make a 50.% ~~volume~~/~~volume~~ solution. If the specific gravity of ethylene glycol is 1.113, calculate a) the boiling and b) the freezing point of this solution. Assume that water is the solvent.
- Using the solution made in problem 3, calculate the a) k_b and b) k_f for ethylene glycol assuming it is the solvent.
- Calculate the molecular mass of ethylene glycol if a 150 g sample is placed into 250 g of ethyl alcohol, which has a boiling point of 78°C and the resulting mixture has a boiling point of 89.8°C .

Colligative Properties Homework Answer

Daniel F McAuley



Colligative Properties Homework Answer:

Basic Physical Pharmacy Joseph K. H. Ma, Boka Hadzija, 2013 Basic Physical Pharmacy provides a thorough yet accessible overview of the principles of physical pharmacy and their application in drug formulation and administration. This definitive guide to physical pharmacy covers all types of pharmaceuticals from traditional forms and dosages to nanotechnology based novel dosage design. Authored by two nationally recognized pharmaceutical scientists and active pharmacy faculty, Basic Physical Pharmacy is clearly organized into four sections: Physical Pharmacy in Solutions, Solid Dosage Forms, Polyphasic Systems, and Drug Delivery and Novel Drug Delivery Systems. Students can build upon their chemistry education to learn the physicochemical properties of drugs and their therapeutic effects on the body. With a highly accessible approach, Basic Physical Pharmacy will help students comprehend and apply the principles of physical pharmacy in clinical practice. Covers major drug products and delivery systems. Features current trends in pharmaceutical research and development including nanotechnology based dosage design. Includes many examples of useful equations and formulation methods. Contains over 200 illustrations, photos, and tables. Topics include: Solutions, Ionization of Drugs in Solutions, Buffers and Buffered Solutions, Drug Solubility, Diffusion and Dissolution, Distribution Phenomena, Complexation and Protein Binding, Interfacial Phenomena, Rheology, Colloids, Suspensions and Emulsions, Semisolid Dosage Forms, Dermatologicals, Suppositories, Powders, Capsules, Tablets, Aerosols, Sterile Dosage Forms, Ophthalmic Formulations, Radiopharmaceuticals, Modified Release Drug Delivery Systems, Biotechnology Products, Drug Product Stability. Each new print textbook includes an access code for the online Companion Website. Ebooks do not include access to the Companion Website. Access to the Companion Website may also be purchased separately under the RESOURCES tab. FOR STUDENTS: Student Companion Website includes Cross Words, Flash Cards, Interactive Glossary, Matching Questions, Instructor Resources, Answers to End of Chapter Questions, Image Bank, Power Point Presentations, Test Bank. **Dr. Salm's Chemistry Problem Drill Book** Robert J. Salm, Robert J. Salm, Robert J. Salm Ph.D., 2005. This drill book contains many common problem types that are asked in General Chemistry classes in High School and College. This work will give you practice with the major problem types as you prepare for finals and standardized tests. **Understanding Chemistry: Chemical systems: thermochemistry, kinetics, and colligative properties**, 1967 **Chemistry Essentials: NEET & JEE Preparation-Vol 2 (XII)** Dr. Manoj Majumder, 2025-03-13

Water Martin Allen, 1998. Reproducible activities correlated to the National Science Education Standards that engage students' minds as they observe, examine, investigate the properties of water from surface tension, density to salinity, buoyancy and much more. **Laboratory Manual to Accompany Chemistry, [by] Stanley R. Radel, Marjorie H. Navidi** Arthur D. Baker, 1990. *Source Book for Chemistry Teachers* William Thomas Lippincott, 1981. [CSIR NET Life Science - Unit 1 - Principles of Biochemistry](#) Mr. Rohit Manglik, 2024-07-02. EduGorilla Publication is a trusted name in the education sector committed to empowering learners with high quality study materials and resources. Specializing in competitive exams and

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Lehninger Principles of Biochemistry David L. Nelson, Michael M. Cox, 2012-11-21 Clear writing and illustrations Clear explanations of difficult concepts Clear communication of the ways in biochemistry is currently understood and practiced For over 35 years in edition after bestselling edition Principles of Biochemistry has put those defining principles into practice guiding students through a coherent introduction to the essentials of biochemistry without overwhelming them

Colligative Properties Lifeliqe, 2019 This lesson plan covers how vapor pressure freezing point and boiling point of a solvent change when a solution is formed **COLLIGATIVE PROPERTIES** NARAYAN CHANGDER, 2024-05-16 Note Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u gmail com You can also get full PDF books in quiz format on our youtube channel <https://www.youtube.com/smartquiziz> I will send you a PDF version of this workbook This book has been designed for candidates preparing for various competitive examinations It contains many objective questions specifically designed for different exams Answer keys are provided at the end of each page It will undoubtedly serve as the best preparation material for aspirants This book is an engaging quiz eBook for all and offers something for everyone This book will satisfy the curiosity of most students while also challenging their trivia skills and introducing them to new information Use this invaluable book to test your subject matter expertise Multiple choice exams are a common assessment method that all prospective candidates must be familiar with in today's academic environment Although the majority of students are accustomed to this MCQ format many are not well versed in it To achieve success in MCQ tests quizzes and trivia challenges one requires test taking techniques and skills in addition to subject knowledge It also provides you with the skills and information you need to achieve a good score in challenging tests or competitive examinations Whether you have studied the subject on your own read for pleasure or completed coursework it will assess your knowledge and prepare you for competitive exams quizzes trivia and more *U.S. Government Research and Development Reports Index* ,1968 **Bibliography of Scientific and Industrial Reports** ,1968 Observing Colligative Properties ,2009-01-01 *Inquiries in Science Chemistry Series Observing Colligative Properties Teacher's Guide*

Properties of Matter: Mixtures and Solutions Gr. 5-8 George Graybill, 2015-09-01 This is the chapter slice Mixtures and Solutions from the full lesson plan Properties of Matter Discover what matter is and is not Learn about and the difference between a mixture and a solution Chocked full with hands on activities to understand the various physical and

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