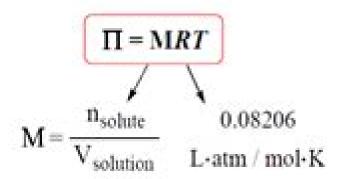
Key Equations for Colligative Properties

Raoult's Law: Relationship between the Vapor Pressure of a Solution (P_{solution}), the Mole Fraction of the Solvent (χ_{solvent}), and the Vapor Pressure of the Pure Solvent (P_{solvent})

The Vapor Pressure of a Solution Containing Two Volatile Components

$$P_{A} = \chi_{A} P^{o}_{A}$$
 $P_{B} = \chi_{B} P^{o}_{B}$
$$P_{total} = P_{A} + P_{B}$$

Relationship between Osmotic Pressure (Π) , Molarity (M), the Ideal Gas Constant (R), and Temperature (T, in K).



The dependance of ΔT_f and ΔT_b on the Molality (m), and Freezing Point Depression Constant (K_f) and Boiling Point Elevation Constant (K_b) respectively.

$$\Delta T_f = m \cdot K_f$$

$$\Delta T_b = m \cdot K_b$$

$$m = \frac{\text{mol (solute)}}{\text{kg (solvent)}}$$

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