



Intermolecular forces

Covalent Compounds Unit Study Guide

Name	
Teacher	
Date:	
Period	

Reference:

Chemistry in the Community, pp 181-183. Naming covalent compounds worksheet Viscosity Lab Exploring Properties of Ionic and Covalent compounds

Vocabulary

Understand and apply the definitions of the following terms:

Covalent Octet rule Valence electrons Polymers. e dot diagrams (Lewis dot diagrams) Viscosity. Boiling points Nonpolar

Targeted Skills

- Diagram a covalent molecule using a e dot diagrams (Lewis dot diagrams).
- Know the following properties:
 - lonic compounds = the transfer of electron(s), i.e. "gain or lose"

Metal cations & nonmetal anion (or polyatomic ion)

Net charge = zero (criss cross charges) X "ayo = X₀Y_a

Dissociate (dissolve) in water

Molecular compounds = sharing electrons by forming a covalent bond

2 or more nonmetals

Prefixes indicate # of each atom

Polar molecules have a 5" & a 5"

Polar molecules dissolve in water

Nonpolar molecules do not have oppositely charged ends, therefore do not dissolve in water.

- Octet rule = atoms share, gain, or lose e- to obtain full outer energy level (usually 8 e-)
- Name covalent compounds using prefixes: (mono, di, tri, tetra, penta, hexa, hepta, octa, nona, deca).
- Compare/contrast the arrangement of atoms in molecules to the arrangement of atoms in ionic compounds.
- Describe the influence of intermolecular forces on physical and chemical properties of covalent compounds. Example: Boiling points, viscosity, and density.
- Be able to describe properties (boiling points, melting points, viscosity, solubility, conductivity) of polymers, ionic crystals, metallic substances, atoms and molecules.

Make sure you remember from past tests?

Atomic structure Significant figures Accuracy & precision Mindurees Nomenclature

Solubility Safety Elements lonic compound formula writing

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Table of Contents Covalent Compounds Study Guide

- 1. Understanding the eBook Covalent Compounds Study Guide
 - The Rise of Digital Reading Covalent Compounds Study Guide
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Covalent Compounds Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Compounds Study Guide
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Covalent Compounds Study Guide
 - Personalized Recommendations
 - Covalent Compounds Study Guide User Reviews and Ratings
 - o Covalent Compounds Study Guide and Bestseller Lists
- 5. Accessing Covalent Compounds Study Guide Free and Paid eBooks
 - Covalent Compounds Study Guide Public Domain eBooks
 - Covalent Compounds Study Guide eBook Subscription Services
 - Covalent Compounds Study Guide Budget-Friendly Options

- 6. Navigating Covalent Compounds Study Guide eBook Formats
 - o ePub, PDF, MOBI, and More
 - Covalent Compounds Study Guide Compatibility with Devices
 - o Covalent Compounds Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Compounds Study Guide
 - Highlighting and Note-Taking Covalent Compounds Study Guide
 - Interactive Elements Covalent Compounds Study Guide
- 8. Staying Engaged with Covalent Compounds Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Compounds Study Guide
- 9. Balancing eBooks and Physical Books Covalent Compounds Study Guide
 - $\circ\,$ Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Covalent Compounds Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Compounds Study Guide
 - Setting Reading Goals Covalent Compounds Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Compounds Study Guide
 - Fact-Checking eBook Content of Covalent Compounds Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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