# Chemical Bonding

# MIXED REVIEW

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

bond energy b. covalent bond enthalpy of vaporization c. metallic bond

2. Use the electronegativity values shown in Figure 20, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

 ionic
 a. H—F
 nonpolar covalent
 d. H—H

 ionic
 b. Na—Cl
 polar covalent
 e. H—C

 polar covalent
 c. H—O
 polar covalent
 f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.
Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

# **Covalent Bonding Study Guide Answers Pearson Chemistry**

**D** Kirk

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#### **Table of Contents Covalent Bonding Study Guide Answers Pearson Chemistry**

- 1. Understanding the eBook Covalent Bonding Study Guide Answers Pearson Chemistry
  - The Rise of Digital Reading Covalent Bonding Study Guide Answers Pearson Chemistry
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Covalent Bonding Study Guide Answers Pearson Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Covalent Bonding Study Guide Answers Pearson Chemistry
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Covalent Bonding Study Guide Answers Pearson Chemistry

- Personalized Recommendations
- Covalent Bonding Study Guide Answers Pearson Chemistry User Reviews and Ratings
- Covalent Bonding Study Guide Answers Pearson Chemistry and Bestseller Lists
- 5. Accessing Covalent Bonding Study Guide Answers Pearson Chemistry Free and Paid eBooks
  - Covalent Bonding Study Guide Answers Pearson Chemistry Public Domain eBooks
  - Covalent Bonding Study Guide Answers Pearson Chemistry eBook Subscription Services
  - Covalent Bonding Study Guide Answers Pearson Chemistry Budget-Friendly Options
- 6. Navigating Covalent Bonding Study Guide Answers Pearson Chemistry eBook Formats
  - o ePub, PDF, MOBI, and More
  - Covalent Bonding Study Guide Answers Pearson Chemistry Compatibility with Devices
  - Covalent Bonding Study Guide Answers Pearson Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Covalent Bonding Study Guide Answers Pearson Chemistry
  - Highlighting and Note-Taking Covalent Bonding Study Guide Answers Pearson Chemistry
  - Interactive Elements Covalent Bonding Study Guide Answers Pearson Chemistry
- 8. Staying Engaged with Covalent Bonding Study Guide Answers Pearson Chemistry
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Covalent Bonding Study Guide Answers Pearson Chemistry
- 9. Balancing eBooks and Physical Books Covalent Bonding Study Guide Answers Pearson Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Covalent Bonding Study Guide Answers Pearson Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Bonding Study Guide Answers Pearson Chemistry
  - Setting Reading Goals Covalent Bonding Study Guide Answers Pearson Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Bonding Study Guide Answers Pearson Chemistry

- Fact-Checking eBook Content of Covalent Bonding Study Guide Answers Pearson Chemistry
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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