

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy _____	a. ionic bond
_____ bond energy _____	b. covalent bond
_____ enthalpy of vaporization _____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic _____	a. H—F	_____ nonpolar covalent _____	d. H—H
_____ ionic _____	b. Na—Cl	_____ polar covalent _____	e. H—C
_____ polar covalent _____	c. H—O	_____ polar covalent _____	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

# Covalent Bonding 8 1 Study Guide Answer Key

**Arshad Iqbal**



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### **Table of Contents Covalent Bonding 8 1 Study Guide Answer Key**

1. Understanding the eBook Covalent Bonding 8 1 Study Guide Answer Key
  - The Rise of Digital Reading Covalent Bonding 8 1 Study Guide Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding 8 1 Study Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Covalent Bonding 8 1 Study Guide Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding 8 1 Study Guide Answer Key
  - Personalized Recommendations

- Covalent Bonding 8 1 Study Guide Answer Key User Reviews and Ratings
- Covalent Bonding 8 1 Study Guide Answer Key and Bestseller Lists
- 5. Accessing Covalent Bonding 8 1 Study Guide Answer Key Free and Paid eBooks
  - Covalent Bonding 8 1 Study Guide Answer Key Public Domain eBooks
  - Covalent Bonding 8 1 Study Guide Answer Key eBook Subscription Services
  - Covalent Bonding 8 1 Study Guide Answer Key Budget-Friendly Options
- 6. Navigating Covalent Bonding 8 1 Study Guide Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Covalent Bonding 8 1 Study Guide Answer Key Compatibility with Devices
  - Covalent Bonding 8 1 Study Guide Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Covalent Bonding 8 1 Study Guide Answer Key
  - Highlighting and Note-Taking Covalent Bonding 8 1 Study Guide Answer Key
  - Interactive Elements Covalent Bonding 8 1 Study Guide Answer Key
- 8. Staying Engaged with Covalent Bonding 8 1 Study Guide Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Covalent Bonding 8 1 Study Guide Answer Key
- 9. Balancing eBooks and Physical Books Covalent Bonding 8 1 Study Guide Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Covalent Bonding 8 1 Study Guide Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Bonding 8 1 Study Guide Answer Key
  - Setting Reading Goals Covalent Bonding 8 1 Study Guide Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Bonding 8 1 Study Guide Answer Key
  - Fact-Checking eBook Content of Covalent Bonding 8 1 Study Guide Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
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