



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Draw a molecular formula for each of the following compounds.

- a. Silicon tetrachloride b. Nitrogen dioxide c. Hydrogen sulfide
d. Carbon disulfide e. Sulfur dioxide

2. Write the formula for the following molecular compounds.

- a. Carbon monoxide b. Dinitrogen pentoxide c. Hydrogen cyanide
d. Nitrogen dioxide e. Sulfur trioxide

3. Draw the Lewis structure for each of the following molecular compounds.
 a. Nitrogen dioxide (Note: Nitrogen dioxide has an unpaired electron and is a free radical.)
 b. Hydrogen cyanide (Note: Carbon has a lone pair of electrons and a negative charge.)
 c. Hydrogen peroxide (Note: Each oxygen has two lone pairs of electrons.)

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF.



2. Draw the Lewis structure for methane, CH₄.



3. Draw the Lewis structure for carbon dioxide, CO₂.



4. Draw the Lewis structure for acetylene, C₂H₂.

5. Draw the Lewis structure for each of the following compounds. Be sure to show all lone pairs of electrons.



(Note: The structure above is a skeletal structure of a branched alkane.)

SECTION 8.3 BONDING: THE ORBITAL

1. Draw the orbital diagram for each of the following molecules. Be sure to show all lone pairs of electrons.
 a. Nitrogen dioxide (Note: Nitrogen dioxide has an unpaired electron and is a free radical.)
 b. Hydrogen cyanide (Note: Carbon has a lone pair of electrons and a negative charge.)

Covalent Bonding Practice Problems Chapter 8

Paul B. Kelter



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