

REPORT SHEET

Compounds and Their Bonds

LAB

9

A. Ions: Transfer of Electrons

Element	Atomic #	1. Electron Configuration of Atom	2. Electron Dot Symbol	3. Loss or Gain	4. Electron Configuration of Ion	5. Ionic Charge	6. Symbol of Ion	7. Name of Ion
Sodium	11	$1s^2 2s^2 2p^6 3s^1$	Na [•]	lose 1 e ⁻	$1s^2 2s^2 2p^6$	1 ⁺	Na ⁺	Sodium ion
Nitrogen	7	$1s^2 2s^2 2p^3$	• N •	Gain 3 e ⁻	$1s^2 2s^2 2p^6$	3 ⁻	N ³⁻	Nitride ion
Aluminum	13	$1s^2 2s^2 2p^6 3s^2 3p^1$	• Al •	Lose 3 e ⁻	$1s^2 2s^2 2p^6$	3 ⁺	Al ³⁺	Aluminum ion
Chlorine	17	$1s^2 2s^2 2p^6 3s^2 3p^5$: Cl •	Gain 1 e ⁻	$1s^2 2s^2 2p^6 3s^2 3p^6$	1 ⁻	Cl ⁻	Chloride ion
Calcium	20	$1s^2 2s^2 2p^6 3s^2 3p^4 4s^2$	Ca ^{••}	Lose 2 e ⁻	$1s^2 2s^2 2p^6 3s^2 3p^4$	2 ⁺	Ca ²⁺	Calcium ion
Oxygen	8	$1s^2 2s^2 2p^4$	• O •	Gain 2 e ⁻	$1s^2 2s^2 2p^6$	2 ⁻	O ²⁻	Oxide ion

B. Ionic Compounds and Formulas

1. Physical properties

Compound	Appearance	Density	Melting Point
Sodium chloride, NaCl	White, solid, crystalline	2.16 g/cm ³	801°C

2. Formulas of ionic compounds

Name	Positive Ion	Negative Ion	Formula
Lithium iodide	Li ¹⁺	I ¹⁻	LiI
Aluminum oxide	Al ³⁺	O ²⁻	Al ₂ O ₃
Calcium sulfide	Ca ²⁺	S ²⁻	CaS
Magnesium bromide	Mg ²⁺	Br ¹⁻	MgBr ₂
Potassium nitride	K ¹⁺	N ³⁻	K ₃ N
Sodium fluoride	Na ¹⁺	F ¹⁻	NaF

3. Names of ionic compounds

Formula	Positive Ion	Negative Ion	Name
K ₂ S	K ¹⁺	S ²⁻	Potassium sulfide
BaF ₂	Ba ²⁺	F ¹⁻	Barium fluoride
MgO	Mg ²⁺	O ²⁻	Magnesium oxide
Na ₃ N	Na ¹⁺	N ³⁻	Sodium nitride
AlCl ₃	Al ³⁺	Cl ¹⁻	Aluminum chloride
Mg ₃ P ₂	Mg ²⁺	P ³⁻	Magnesium phosphide

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