

Name Answer Key Pd _____ Date _____

Accelerated Chemistry: Average Atomic Mass Worksheet

Calculate the average atomic mass for each element based on the natural abundance of its isotopes.

1. Find the average atomic mass for Li if 7.5% of Li atoms are ^6Li with a mass of 6.0151223 amu and 92.5% are ^7Li with a mass of 7.0160041 amu.

$$(6.0151223 \text{ amu}) 0.075 + (7.0160041 \text{ amu}) 0.925 = 6.9 \text{ amu}$$

2. Find the average atomic mass for B if 19.9% of B atoms are ^{10}B with a mass of 10.0129371 amu and 80.1% are ^{11}B with a mass of 11.0093055 amu.

$$(10.0129371 \text{ amu}) 0.199 + (11.0093055 \text{ amu}) 0.801 = 10.8 \text{ amu}$$

3. Find the average atomic mass for Cl if 75.78% of Cl atoms are ^{35}Cl with a mass of 34.96885271 amu and 24.22% are ^{37}Cl with a mass of 36.96590260 amu.

$$(34.96885271 \text{ amu}) 0.7578 + (36.96590260 \text{ amu}) 0.2422 = 35.45 \text{ amu}$$

4. Find the average atomic mass for Mg if 78.99% of Mg atoms are ^{24}Mg with a mass of 23.9850419 amu, 10.00% are ^{25}Mg with a mass of 24.9858370 amu, and 11.01% are ^{26}Mg with a mass of 25.9825930 amu.

$$(23.9850419 \text{ amu}) 0.7899 + (24.9858370 \text{ amu}) 0.1000 + (25.9825930 \text{ amu}) 0.1101 = 24.31 \text{ amu}$$

5. There are 2 isotopes of copper that occur naturally; ^{63}Cu and ^{65}Cu . The ^{63}Cu atoms have a mass of 62.929601 amu and the ^{65}Cu atoms have a mass of 64.927794 amu. What is the percent natural abundance for each isotope? $^{63}\text{Cu} = 69.15\%$ $^{65}\text{Cu} = 30.85\%$

$$63.546 \text{ amu} = (62.929601 \text{ amu})x + (64.927794 \text{ amu})(1-x)$$

the 2 isotope abundances must equal all of the copper so
let $^{63}\text{Cu} = x$ and $^{65}\text{Cu} = 1-x$. Now isolate and solve for x,

$$63.546 \text{ amu} = (62.929601 \text{ amu})x + (64.927794 \text{ amu}) - (64.927794 \text{ amu})x$$

$$63.546 \text{ amu} - 64.927794 \text{ amu} = (62.929601 \text{ amu} - 64.927794 \text{ amu})x = -1.998193x$$

$$x = \frac{-1.382}{-1.998193} = 0.6915 \stackrel{63}{\text{Cu}}; \quad ^{65}\text{Cu} = 1 - 0.6915 = 0.3085$$

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