

Balancing Equations Worksheet

Double replacement reaction

- 1) $\underline{\hspace{1cm}} \text{Na}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{KOH} \rightarrow \underline{\hspace{1cm}} \text{NaOH} + \underline{\hspace{1cm}} \text{K}_3\text{PO}_4$ *Na PO₄*
- 2) $\underline{\hspace{1cm}} \text{MgF}_2 + \underline{\hspace{1cm}} \text{Li}_2\text{CO}_3 \rightarrow \underline{\hspace{1cm}} \text{MgCO}_3 + \underline{\hspace{1cm}} \text{LiF}$
- 3) $\underline{\hspace{1cm}} \text{P}_4 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{P}_2\text{O}_3$
- 4) $\underline{\hspace{1cm}} \text{RbNO}_3 + \underline{\hspace{1cm}} \text{BeF}_2 \rightarrow \underline{\hspace{1cm}} \text{Be}(\text{NO}_3)_2 + \underline{\hspace{1cm}} \text{RbF}$
- 5) $\underline{\hspace{1cm}} \text{AgNO}_3 + \underline{\hspace{1cm}} \text{Cu} \rightarrow \underline{\hspace{1cm}} \text{Cu}(\text{NO}_3)_2 + \underline{\hspace{1cm}} \text{Ag}$
- 6) $\underline{\hspace{1cm}} \text{CF}_4 + \underline{\hspace{1cm}} \text{Br}_2 \rightarrow \underline{\hspace{1cm}} \text{CBr}_4 + \underline{\hspace{1cm}} \text{F}_2$
- 7) $\underline{\hspace{1cm}} \text{HCN} + \underline{\hspace{1cm}} \text{CuSO}_4 \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{SO}_4 + \underline{\hspace{1cm}} \text{Cu}(\text{CN})_2$
- 8) $\underline{\hspace{1cm}} \text{GaF}_3 + \underline{\hspace{1cm}} \text{Cs} \rightarrow \underline{\hspace{1cm}} \text{CsF} + \underline{\hspace{1cm}} \text{Ga}$
- 9) $\underline{\hspace{1cm}} \text{BaS} + \underline{\hspace{1cm}} \text{PtF}_2 \rightarrow \underline{\hspace{1cm}} \text{BaF}_2 + \underline{\hspace{1cm}} \text{PtS}$
- 10) $\underline{\hspace{1cm}} \text{N}_2 + \underline{\hspace{1cm}} \text{H}_2 \rightarrow \underline{\hspace{1cm}} \text{NH}_3$
- 11) $\underline{\hspace{1cm}} \text{NaF} + \underline{\hspace{1cm}} \text{Br}_2 \rightarrow \underline{\hspace{1cm}} \text{NaBr} + \underline{\hspace{1cm}} \text{F}_2$
- 12) $\underline{\hspace{1cm}} \text{Pb}(\text{OH})_2 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{O} + \underline{\hspace{1cm}} \text{PbCl}_2$
- 13) $\underline{\hspace{1cm}} \text{AlBr}_3 + \underline{\hspace{1cm}} \text{K}_2\text{SO}_4 \rightarrow \underline{\hspace{1cm}} \text{KBr} + \underline{\hspace{1cm}} \text{Al}_2(\text{SO}_4)_3$
- 14) $\underline{\hspace{1cm}} \text{CH}_4 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$

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